



X App Compare Tutorial

William James Sutherland



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 $3 \text{ hbr} + 1 \text{ khco}_3 \rightarrow 3 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2$
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 $2 \text{ c}_6\text{h}_{10} + 17 \text{ o}_2 \rightarrow 12 \text{ co}_2 + 10 \text{ h}_2\text{o}$
 $3 \text{ hbr} + 1 \text{ khco}_3 \rightarrow 3 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2$
using the following equation
 $2 \text{ naoh} + \text{ h}_2\text{so}_4 \rightarrow 2 \text{ h}_2\text{o} + \text{ na}_2\text{so}_4$

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 $\text{ce} + \text{zn} + \text{cucl}_2 \rightarrow \text{zncl}_2 + \text{cu} + \text{zn} + \text{cucl}_x$
2 moles of ce will be produced from 23.0 g of ce and 23.0 g of zn assuming ce and cucl₂ are available in excess

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 $3 \text{ hbr} + 1 \text{ khco}_3 \rightarrow 3 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2$
using the following equation
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