

7-2 Review and Reinforcement

Covalent Bonding

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line provided.

- | | |
|---|---|
| ___ covalent
___ true
___ 4 shared (2 pairs)
___ true
___ structural
___ lone/non-bonding
___ equally
___ triple | 1. A group of atoms united by <u>ionic</u> bonds is called a molecule.
2. A covalent bond is formed by a <u>shared</u> pair of electrons.
3. A double covalent bond consists of <u>two</u> shared electrons.
4. A <u>molecular</u> formula tells you how many atoms are in a single molecule of the compound.
5. The <u>empirical</u> formula for a molecule specifies which atoms are bonded to each other in the molecule.
6. A pair of electrons not involved in bonding is called a <u>shared electron pair</u> .
7. In nonpolar covalent bonds, the electrons are shared <u>unequally</u> between two atoms.
8. Lewis structures use a triple dash to represent a <u>double</u> bond. |
|---|---|

Use the information in Figure 7-19 of the textbook to calculate the electronegativity differences between each pair of elements listed in the chart below. Then indicate whether the bond would be nonpolar covalent, polar covalent, or ionic. Write your answers in the chart.

		Electronegativity Difference	Type of Bond
9.	H—O	1.4	polar covalent
10.	C—H	.4	non-polar covalent
11.	K—F	3.2	ionic
12.	N—H	.9	polar covalent
13.	Na—F	3.1	ionic
14.	O—Cl	.5	polar covalent

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Thomas Griffiths



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Table of Contents 7 2 Review And Reinforcement Answer Key

1. Understanding the eBook 7 2 Review And Reinforcement Answer Key
 - The Rise of Digital Reading 7 2 Review And Reinforcement Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 7 2 Review And Reinforcement Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 7 2 Review And Reinforcement Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 7 2 Review And Reinforcement Answer Key
 - Personalized Recommendations
 - 7 2 Review And Reinforcement Answer Key User Reviews and Ratings
 - 7 2 Review And Reinforcement Answer Key and Bestseller Lists
5. Accessing 7 2 Review And Reinforcement Answer Key Free and Paid eBooks

- 7 2 Review And Reinforcement Answer Key Public Domain eBooks
 - 7 2 Review And Reinforcement Answer Key eBook Subscription Services
 - 7 2 Review And Reinforcement Answer Key Budget-Friendly Options
6. Navigating 7 2 Review And Reinforcement Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 7 2 Review And Reinforcement Answer Key Compatibility with Devices
 - 7 2 Review And Reinforcement Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 7 2 Review And Reinforcement Answer Key
 - Highlighting and Note-Taking 7 2 Review And Reinforcement Answer Key
 - Interactive Elements 7 2 Review And Reinforcement Answer Key
 8. Staying Engaged with 7 2 Review And Reinforcement Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 7 2 Review And Reinforcement Answer Key
 9. Balancing eBooks and Physical Books 7 2 Review And Reinforcement Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 7 2 Review And Reinforcement Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 7 2 Review And Reinforcement Answer Key
 - Setting Reading Goals 7 2 Review And Reinforcement Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 7 2 Review And Reinforcement Answer Key
 - Fact-Checking eBook Content of 7 2 Review And Reinforcement Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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