

 business

# Reddit for Business



# Reddit Pro Guide

**Michal Rosen-Zvi**



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 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$   
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$   
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$   
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$   
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4$   
 $3\text{CH}_4 + 8\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

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web oct 2 2019 let's start by balancing the oxygen atoms to do this make the oxygen atoms as six on either side of the chemical equation  $2\text{Fe}_2\text{O}_3 + \text{C} \rightarrow 3\text{CO}_2$  on towards the next step now step 3 once you are done with this check out if the number of atoms for the other elements is also equal on both the sides

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web 1 78 348 stoichiometry helps us determine how much substance is needed or is present things that can be measured are reactants and products mass molecular weight chemical equations formulas stoichiometric coefficient the stoichiometric coefficient or stoichiometric number is the number of molecules that participate in the reaction

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web expert answer transcribed image text balancing equations and simple stoichiometry balance the following equations  
 $1\text{ N}_2 + \text{F}_2 \rightarrow 2\text{NF}_3$   
 $2\text{C}_6\text{H}_{10}\text{O}_2 + \text{CO}_2 + \text{H}_2\text{O} \rightarrow 3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 + \text{GaBr} + \text{NaBr} + \text{SnO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{NF}_3 + 3\text{SnF}_2 + \text{N}_2\text{O}_3$  using the equation from problem 2 above answer the following

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web learning objectives to balance equations that describe reactions in solution to calculate the quantities of compounds produced or consumed in a chemical reaction to solve quantitative problems involving the stoichiometry of reactions in solution

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 $1\text{ N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$   
 $2\text{C}_6\text{H}_{10}\text{O}_2 + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$   
 $3\text{HBr} + 1\text{KHCO}_3 + 1\text{H}_2\text{O} + 1\text{KBr} + \text{CO}_2 + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{SO}_3 + 6\text{NaBr} + 3\text{SnO}_2 + 3\text{NF}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$  using the following equation  $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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web jul 22 2013 balance the following equations  
 $\text{br}_2 + \text{n}_2 \rightarrow \text{br}_3\text{n}_2$   
 $\text{c}_6\text{h}_{10} + \text{o}_2 \rightarrow \text{co}_2 + \text{h}_2\text{o}$   
 $\text{h}_2\text{o} + \text{kbr} + \text{co}_2 \rightarrow \text{gabr}_3 + \text{na}_2\text{so}_3 + \text{ga}_2\text{so}_3 + \text{nabr}$   
 $\text{br}_2 + \text{sno}_2 \rightarrow \text{snf}_2 + \text{n}_2\text{o}_3$   
 $\text{naoh} + \text{h}_2\text{so}_4 \rightarrow \text{h}_2\text{o} + \text{na}_2\text{so}_4$

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web the stoichiometry of a balanced chemical equation identifies the maximum amount of product that can be obtained the stoichiometry of a reaction describes the relative amounts of reactants and products in a balanced chemical equation

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 $\text{c}_6\text{h}_{10} + \text{o}_2 \rightarrow \text{co}_2 + \text{h}_2\text{o}$   
 $\text{hbr} + \text{khco}_3 \rightarrow \text{h}_2\text{o} + \text{kbr} + \text{co}_2$   
 $\text{gabr} + \text{na}_2\text{so}_3 \rightarrow \text{gaz} + \text{so}_3 + \text{nabr} + \text{sno} + \text{nf}_3 + \text{snf}_2$

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web may 28 2020 write the balanced equation for the reaction and calculate how much  $\text{si}_2$  is required to produce 3.00 kg of  $\text{sic}$   
 $\text{ce} + \text{si}_2 \rightarrow \text{sic}$

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 $\text{n}_2 + \text{f}_2 \rightarrow \text{nf}_3$   
 $\text{c}_6\text{h}_{10} + \text{o}_2 \rightarrow \text{co}_2 + \text{h}_2\text{o}$   
 $\text{hbr} + \text{khco}_3 \rightarrow \text{h}_2\text{o} + \text{kbr} + \text{co}_2$   
 $\text{gabr} + \text{na}_2\text{so}_3 \rightarrow \text{gaz} + \text{so}_3 + \text{sno} + \text{nf}_3 + \text{snf}_2 + \text{n}_2\text{o}_3$   
 $\text{nabr}$  solve the following stoichiometry grams problems 6

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web gabrielle m 9 years ago i m working on chemical reactions double and single replacement on flvs now my first question for this video is how do you have 4 aluminum atoms when it says  $2\text{Al}$  subscript 2 do you just add the 2 and the subscript 2 same thing goes with the  $\text{O}_3$  i am so lost

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web google classroom balance the following chemical equation  $\text{Mg(OH)}_2 + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2\text{O}$  note all reactants and products require a coefficient of at least one stuck review related articles

videos or use a hint report a problem  $7.41 \times x y \theta \pi$  8 5 2 0 9 6

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