



Student Exploration: Balancing Chemical Equations

(This resource is intended for use with the Student Exploration: Balancing Chemical Equations.)

Essential Question: How do you balance a chemical equation? What are the coefficients in a balanced chemical equation? How do you use a balanced chemical equation to predict the amount of product that will be formed?

Prior Knowledge: Students should understand the following concepts:

1. Atoms, mass of atoms, molar mass, and grams of sugar. Students may not know that 1 mole of a substance is 6.022 × 10²³ particles of that substance.
2. It requires the same amount of mass to make a product. For example, if you have 10 grams of reactants, you will have 10 grams of products. All atoms of matter are conserved in a reaction.

Student Worksheet

Read the introduction and answer the questions. Then write a balanced chemical equation for the reaction.

Introduction: In chemistry, the amount of reactants required to produce a product is called a balanced chemical equation. The amount of product that will be formed depends on the amount of reactants used.

Example: Consider the reaction of sodium



1. Look at the **balanced** equation above. It is balanced because the number of atoms of each element is the same on both sides. The coefficients tell you how many molecules of each substance are involved in the reaction. For the number of each substance within the equation:

Reactants	Na	Cl ₂	Products	NaCl
2	2	1	2	2

2. It is **balanced** because the number of each kind of particle (atom or molecule) is the same on both sides. The number of particles (atoms or molecules) of each substance is:

Reactants	2Na	1Cl ₂	Products	2NaCl
Atoms	2Na	2Cl	2Na	2Cl

How many atoms of each element are there on each side of the equation? How do you know it is balanced?

Introduction

Read the introduction and answer the questions.

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Example

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How many atoms of each element are there on each side of the equation? How do you know it is balanced?



Explore Learning Stoichiometry Answers

Thomas Griffiths



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