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Mike Meyers



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calorimeter with a total heat capacity of 15 48 kJ $^{\circ}\text{C}^{-1}$ the temperature of the calorimeter and its contents changes from 19 77 $^{\circ}\text{C}$
to 33 90 $^{\circ}\text{C}$ calculate ΔH for the following reaction as written $2 \text{NH}_3(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g}) + 2 \text{CO}_2(\text{g})$

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final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the
final temperature is closer to t_2 than to t_1 c

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water is 4 184 J $\text{g}^{-1} \text{ }^{\circ}\text{C}^{-1}$ $m = 5 650 \text{ g}$ H_2O 4 184 J $\text{g}^{-1} \text{ }^{\circ}\text{C}^{-1}$ 5 4 $^{\circ}\text{C}$ $q = 1310 \text{ kJ}$

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ice at 34 $^{\circ}\text{C}$ to steam at 138 $^{\circ}\text{C}$ a sample of water absorbs $1 41 \times 10^5 \text{ J}$ find the mass of the sample answers 1 2 68 $\times 10^4 \text{ J}$ 2 6
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