

# Soluzioni

## CAPITOLO 15

### VERIFICA LE TUE CONOSCENZE

#### VALENZA E NUMERO DI OSSIDAZIONE

- 1**
- La valenza di un elemento è il numero di elettroni che l'atomo cede, acquista o condivide quando si lega ad altri atomi.
  - La valenza massima corrisponde al numero romano che contraddistingue il gruppo, quindi, nel caso di un elemento del terzo gruppo della tavola periodica, è 3.
  - Un atomo di carbonio può formare al massimo 4 legami.
- 2** (C) di cariche reali o convenzionali associato a ciascun atomo del composto.
- 3** Regole per il calcolo del numero di ossidazione non corrette [tra parentesi, la versione corretta]
- l'ossigeno ha sempre n.o.  $-2$ ; [l'ossigeno nei composti ha n.o.  $-2$ , tranne nei perossidi, nei superossidi e nel difluoruro di ossigeno, dove ha rispettivamente n.o.  $-1$ ,  $-1/2$ ,  $+2$ ];
  - l'idrogeno ha n.o.  $+1$  nei composti e  $+2$  nella molecola  $H_2$ ; [l'idrogeno nei composti ha n.o.  $+1$  quando è legato a un non metallo e  $-1$  quando è legato a un metallo].

#### SCRIVERE LE FORMULE DEI COMPOSTI A PARTIRE DAI NUMERI DI OSSIDAZIONE

- 4**  $Na_2O$ : in questo modo la somma algebrica dei n.o. è uguale a zero:  $2 \cdot (+1) + (-2) = 0$ .
- 5**  $Ca_3(PO_4)_2$
- 6**
- $SO_4^{2-}$ ; carica dello ione  $= -2$
  - $ClO_7^-$ ; carica dello ione  $= -1$
- 7** La formula  $Cl_2Ba$  è sbagliata, in quanto si scrive a sinistra l'atomo con n.o. positivo ( $Ba_2Cl_2$ ).
- 8** È corretta soltanto: d.  $SO_2$  [tra parentesi, le formule corrette]
- HS [ $H_2S$ ]
  - $Fe_3S_2$  [ $Fe_2S_3$ ]
  - $LiS_2$  [ $Li_2S$ ]

#### LA NOMENCLATURA DEI COMPOSTI BINARI

- 9** Un composto binario è formato da due elementi diversi. I composti binari si dividono in:
- composti contenenti ossigeno: ossidi (es.  $ZnO$ ), perossidi (es.  $H_2O_2$ ), superossidi;
  - composti contenenti idrogeno: idruri (es.  $NaH$ ), idracidi (es.  $HF$ );
  - sali binari (es.  $NaCl$ ,  $CaCl_2$ ).
- 10** In molti composti binari si ha la desinenza *-ure*; tale desinenza non compare negli ossidi.
- 11** La notazione di Stock si utilizza quando gli elementi che formano un composto hanno più di un numero di ossidazione. Comporta l'eliminazione dei prefissi numerici e l'introduzione del numero di ossidazione dell'elemento, come numero romano tra parentesi posto dopo il nome dell'elemento. Si ritiene che in questo modo il nome che contraddistingue l'elemento risulti più breve e più chiaro.
- (n.o. Fe =  $+2$ )     $FeBr_2$     bromuro di ferro(II)  
(n.o. Fe =  $+3$ )     $FeBr_3$     bromuro di ferro(III)
- 12** Lo ione ossido ( $O^{2-}$ ) è formato da un atomo di O con n.o.  $= -2$ .  
Lo ione perossido ( $O_2^{2-}$ ) è formato da due atomi di O entrambi con n.o.  $= -1$ , legati da legame covalente.  
Lo ione superossido ( $O_2^-$ ) è formato da due atomi di O con n.o. medio  $= -1/2$ , legati da legame covalente.  
Un ossido acido è formato da un non metallo e ossigeno, un ossido basico è formato da un metallo e ossigeno.
- 13**
- Idrogeno e metallo: idruri ( $NaH$ ,  $LiH$ );
  - idrogeno e non metallo dei gruppi 16 e 17: idracidi ( $HCl$ ,  $HF$ ,  $HBr$ );
  - metallo e non metallo: sali binari ( $NaCl$ ,  $FeCl_2$ ).
- 14** (B) cianidrico.
- 15**  $K_2O$ : ossido di potassio;  
 $SO_2$ : anidride solforosa (n.o. S =  $+4$ ).  
 $K_2O$  si chiama *ossido* perché è un ossido basico (metallo e ossigeno).  
 $SO_2$  si chiama *anidride* perché è un ossido acido (non metallo e ossigeno).

# Soluzioni Libro Network 1

**J Dewey**



## **Soluzioni Libro Network 1:**

## **Soluzioni Libro Network 1** Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the power of words has are more evident than ever. They have the capability to inspire, provoke, and ignite change. Such could be the essence of the book **Soluzioni Libro Network 1**, a literary masterpiece that delves deep in to the significance of words and their effect on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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