

# 9

## Thermodynamics

- Which of the following statements is correct for the spontaneous absorption of a gas? [2014]
  - $\Delta S$  is negative and therefore,  $\Delta H$  should be highly positive
  - $\Delta S$  is negative and therefore,  $\Delta H$  should be highly negative
  - $\Delta S$  is positive and therefore,  $\Delta H$  should be negative
  - $\Delta S$  is positive and therefore,  $\Delta H$  should also be highly positive
- For the reaction,  $X_2O_4(l) \longrightarrow 2XO_2(g)$ ,  $\Delta U = 2.1 \text{ kcal}$ ,  $\Delta S = 20 \text{ cal K}^{-1}$  at 300 K. Hence,  $\Delta G$  is [2014]
  - 2.7 kcal
  - 2.7 kcal
  - 9.3 kcal
  - 9.3 kcal
- For a given exothermic reaction,  $K_p$  and  $K_p'$  are the equilibrium constants at temperatures  $T_1$  and  $T_2$ , respectively. Assuming that heat of reaction is constant in temperature range between  $T_1$  and  $T_2$ , it is readily observed that [2014]
  - $K_p > K_p'$
  - $K_p < K_p'$
  - $K_p = K_p'$
  - $K_p = \frac{1}{K_p'}$
- A reaction having equal energies of activation for forward and reverse reactions has [2013]
  - $\Delta S = 0$
  - $\Delta G = 0$
  - $\Delta H = 0$
  - $\Delta H = \Delta G = \Delta S = 0$
- In which of the following reactions, standard reaction entropy changes ( $\Delta S^\circ$ ) is positive and standard Gibbs energy change ( $\Delta G^\circ$ ) decreases sharply with increasing temperature? [2012]
  - $C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow CO(g)$
  - $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$
  - $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$
  - $\frac{1}{2}C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow \frac{1}{2}CO_2(g)$
- The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is [2012]
  - 10.52 cal/mol K
  - 21.04 cal/mol K
  - 5.260 cal/mol K
  - 0.526 cal/mol K
- Standard enthalpy of vaporisation  $\Delta_{\text{vap}}H^\circ$  for water at 100°C is 40.66 kJ mol<sup>-1</sup>. The internal energy of vaporisation of water at 100°C (in kJ mol<sup>-1</sup>) is (assume water vapour to behave like an ideal gas). [2012]
  - + 37.56
  - 43.76
  - + 43.76
  - + 40.66
- If the enthalpy change for the transition of liquid water to steam is 30 kJ mol<sup>-1</sup> at 27°C, the entropy change for the process would be [2011]
  - 1.0 J mol<sup>-1</sup>K<sup>-1</sup>
  - 0.1 J mol<sup>-1</sup>K<sup>-1</sup>
  - 100 J mol<sup>-1</sup>K<sup>-1</sup>
  - 10 J mol<sup>-1</sup>K<sup>-1</sup>

# Solution Thermodynamics Important Questions And Answers

**D Keegan**



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