

9/2/16

S. Sai Krishna

Date \_\_\_\_\_  
Page \_\_\_\_\_  
Classmate

## Thermochemistry

$$Q = \Delta E - W$$

$$Q = \Delta E - (-P\Delta V)$$

$$\boxed{Q = \Delta E + P\Delta V}$$

- At constant Volume.

$$Q = \Delta E$$

- At const. Pressure.

$$Q_p = \Delta E + P\Delta V$$

$$\boxed{H = \Delta E + P\Delta V}$$

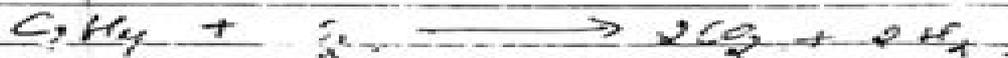
$$\boxed{H = \Delta E + \Delta n RT}$$

$\Delta n =$  no. of moles of Prod  
- no. of moles of react.

→ For  $\Delta H$ .

- If  $\Delta H$  is -ve. reaction is exothermic.
- If  $\Delta H$  is +ve. reaction is endothermic.

Q. The heat of combustion of ethane at 17°C and const volume is  $-322.1 \text{ kJ}$ . Calculate the heat of combustion at constant pressure.



# Engineering Chemistry Full Notes

**RC Schank**



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