

## 5-2 Review and Reinforcement Part II

### Quantum Theory and the Atom

Explain this principle: state the evidence for the discreteness of light (quantization).

photon(s)

Planck constant (symbol)

quantum mechanical model

quantum number(s)

electron density

orbitals

principal energy levels

\_\_\_\_\_ orbital \_\_\_\_\_ 1-4: model to explain around the nucleus where electrons in atoms really found.

\_\_\_\_\_ electron density \_\_\_\_\_ 2: the density of an electron cloud

\_\_\_\_\_ quantum number \_\_\_\_\_ 3: number describing a probability density found in atoms

Planck constant (symbol) is  $h$  where  $h$  is Planck's constant (which sometimes denotes  $2\pi h$ ) and this gives

\_\_\_\_\_ quantum number (symbol) \_\_\_\_\_ 1

\_\_\_\_\_ quantum number (symbol) \_\_\_\_\_ 2: probability probability density by knowing the electron's position and

\_\_\_\_\_ wavelength (symbol)

\_\_\_\_\_ probability density \_\_\_\_\_ 4: the total energy found in atoms

\_\_\_\_\_ electron spin \_\_\_\_\_ 5: the direction of counter-clockwise or clockwise rotation

Explain this evidence: state the nature of the discrete absorption spectrum of hydrogen.

\_\_\_\_\_ 1. \_\_\_\_\_ 1. The discrete lines in hydrogen when the production of ionizing radiation is

- produced
- absorbed
- light is emitted
- absorbed

\_\_\_\_\_ 2. \_\_\_\_\_ 2. The low principal energy level of the hydrogen atom emits only one

- orbital
- of orbitals
- of orbitals
- orbital

# 5 2 Review And Reinforcement Answers

**Marvin L. Bittinger, David Ellenbogen**

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