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web this unit focuses on building an initial understanding of the mole and how to use it to quantify chemicals at a reasonable scale students develop a conceptual understanding of the mole and use molar mass and avogadro s number to calculate the number of atoms or molecules in a substance

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web the mole is an amount unit similar to familiar units like pair dozen gross etc it provides a specific measure of the number of atoms or molecules in a sample of matter one latin connotation for the word mole is large mass or bulk which is consistent with its use as the name for this unit

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web oct 10 2023 mole standard unit $6.02214076 \times 10^{23}$ in chemistry for measuring large quantities of very small entities such as atoms molecules or other specified particles the number of units in a mole also bears the name avogadro s number or avogadro s constant in honor of the italian physicist amedeo avogadro

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web the mole was made the seventh si base unit in 1971 by the 14th cgpm 2019 redefinition of si base units before the 2019 redefinition of the si base units the mole was defined as the amount of substance of a system that contains as many elementary entities as there are atoms in 12 grams of carbon 12 the most common isotope of carbon