

## CHAPTER 7 PRACTICE TEST: ATOMIC STRUCTURE &amp; PERIODIC TRENDS- KEY

1) When a solution of sodium chloride is vaporized in a flame, the color of the flame is:

- A) Blue      **B) Yellow**      C) Green      D) Violet      E) White

2) Electromagnetic radiation travels through vacuum at a speed of \_\_\_\_\_ m/s.

- A) 186,000      **C)  $3.00 \times 10^8$**       E) It depends on  $\lambda$   
 B) 125      D) 10,000       $\lambda$

3) The wavelength of light that has a frequency of  $1.20 \times 10^{13} \text{ s}^{-1}$  is \_\_\_\_\_ m.

- A) 25.0      **B)  $2.50 \times 10^{-2}$**       C) 0.0400      E) 2.5  
 D) 12.0

GIVEN:  $\nu = 1.20 \times 10^{13} \text{ s}^{-1}$

UNKNOWN:  $\lambda$

CONNECTION:  $c = \lambda \nu$

SUBSTITUTE & SOLVE:  $\lambda = c/\nu$

$$\lambda = \frac{3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}}{1.20 \times 10^{13} \text{ s}^{-1}}$$

$$\lambda = 2.50 \times 10^{-2} \text{ m}$$

4) Ham radio operators often broadcast on the 6-meter band. The frequency of this EM radiation is \_\_\_\_\_ MHz.

- A) 500      B) 200      **C) 50**      D) 20      E) 2.0

GIVEN:  $\lambda = 6 \text{ m}$

UNKNOWN:  $\nu$  (Must convert answer to MHz!)

CONNECTION:  $c = \lambda \nu$

SUBSTITUTE & SOLVE:  $\nu = c/\lambda$

$$\nu = \frac{3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}}{6 \text{ m}}$$

$$\nu = 0.5 \times 10^8 \text{ Hz} \times \frac{1 \text{ MHz}}{10^6 \text{ Hz}} = 50 \text{ MHz}$$

5) The energy of a photon of light is \_\_\_\_\_ proportional to its frequency and \_\_\_\_\_ proportional to its wavelength.

- A) directly, directly      C) inversely, directly      E) indirectly, not  
 B) inversely, inversely      **D) directly, inversely**

6) The energy of a photon that has a wavelength of 12.3 nm is \_\_\_\_\_ J.

- A)  $1.51 \times 10^{-17}$       B)  $4.42 \times 10^{-23}$       C)  $1.99 \times 10^{-23}$       D)  $2.72 \times 10^{-20}$       **E)  $1.62 \times 10^{-17}$**

7) The wavelength of a photon that has an energy of  $5.25 \times 10^{-19} \text{ J}$  is \_\_\_\_\_ m.

- A)  $3.79 \times 10^{-7}$**       B)  $2.64 \times 10^6$       C)  $2.38 \times 10^{23}$       D)  $4.21 \times 10^{-24}$       E)  $3.79 \times 10^7$

8) A mole of red photons of wavelength 725 nm has \_\_\_\_\_ kJ of energy.

- A)  $2.74 \times 10^{-29}$       B)  $4.56 \times 10^{-26}$       C)  $6.05 \times 10^{-7}$       **D) 165**      E) 227

GIVENS:  $\lambda = 725 \text{ nm}$  (must convert to m before using formulas)

$$\lambda = 725 \text{ nm} \times \frac{1 \text{ m}}{10^9 \text{ nm}} = 7.25 \times 10^{-7} \text{ m}$$

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## Table of Contents Ap Chemistry Chapter 13 Test

1. Understanding the eBook Ap Chemistry Chapter 13 Test
  - The Rise of Digital Reading Ap Chemistry Chapter 13 Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Chapter 13 Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ap Chemistry Chapter 13 Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Chapter 13 Test
  - Personalized Recommendations
  - Ap Chemistry Chapter 13 Test User Reviews and Ratings
  - Ap Chemistry Chapter 13 Test and Bestseller Lists
5. Accessing Ap Chemistry Chapter 13 Test Free and Paid eBooks
  - Ap Chemistry Chapter 13 Test Public Domain eBooks
  - Ap Chemistry Chapter 13 Test eBook Subscription Services
  - Ap Chemistry Chapter 13 Test Budget-Friendly Options
6. Navigating Ap Chemistry Chapter 13 Test eBook Formats
  - ePub, PDF, MOBI, and More
  - Ap Chemistry Chapter 13 Test Compatibility with Devices
  - Ap Chemistry Chapter 13 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ap Chemistry Chapter 13 Test
  - Highlighting and Note-Taking Ap Chemistry Chapter 13 Test
  - Interactive Elements Ap Chemistry Chapter 13 Test
8. Staying Engaged with Ap Chemistry Chapter 13 Test

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ap Chemistry Chapter 13 Test
9. Balancing eBooks and Physical Books Ap Chemistry Chapter 13 Test
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ap Chemistry Chapter 13 Test
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ap Chemistry Chapter 13 Test
- Setting Reading Goals Ap Chemistry Chapter 13 Test
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chemistry Chapter 13 Test
- Fact-Checking eBook Content of Ap Chemistry Chapter 13 Test
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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