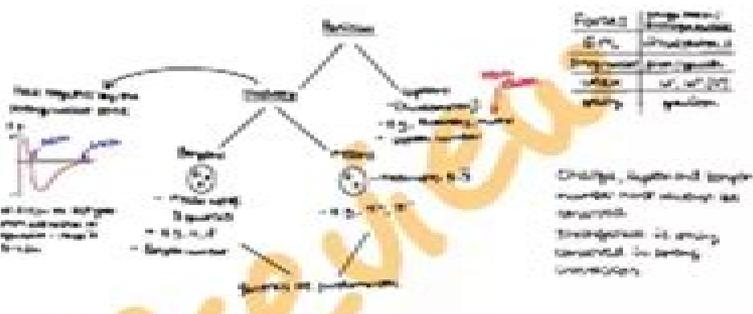


Particles and Quantum



To convert mass to energy... $E=mc^2$
 Annihilation - pair of particles and its corresponding antiparticle collide and their rest energy is converted into photons

Pair production - when photon turns into a particle and its corresponding antiparticle. The photon must have at least the sum amount of energy which was energy of the particles

Excitation
 Exciting means they give electrons enough energy to climb their atomic shells

Alpha particles - Helium nucleus, ${}^4_2\text{He}$
 - Highly ionising, so it is dangerous
 - Relatively penetrating - stopped by a piece of paper or foil on their skin

Beta - fast moving electron $\rightarrow e^-$
 - Medium penetrating and relatively ionising - stopped by a few cm of aluminium
 - Relatively penetrating - stopped by a few cm of aluminium



Gamma - high energy EM ray or photon emitted from a high energy nucleus $\rightarrow \gamma$
 - Highly penetrating - stopped by a few m of concrete

Isotopes - same element - same number of protons
 - different relative atomic mass - different number of neutrons

Nuclear charge - $\frac{Q}{A} [100\%]$
 - charge to mass ratio

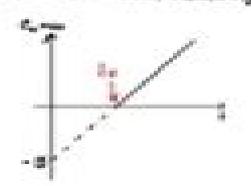
Radioactive - change of a nucleus when spontaneous through a bit of time and
 - 10^{-10} s
 - 100



Photoelectric effect

Use photons of sufficient energy on electrons on the surface of the metal, liberating them

- 1. Photon absorbed - $h\nu$ (total energy)
- 2. Frequency
- 3. Work function - minimum energy to liberate electrons
- 4. Kinetic energy left over after $h\nu$ is subtracted
- 5. Stopped potential - minimum potential needed for electrons to be liberated



The photoelectric effect proved the particle nature of light due to the fact that the maximum kinetic energy of electrons K_{max} is not linearly dependent on the wavelength of light. Instead, it is linearly dependent on the frequency of light. K_{max} is zero for frequencies below the threshold frequency f_0 and increases linearly for frequencies above f_0 .

Stopping potential - the pot. needed to stop any electrons from reaching the other side

- 1. $E_{photon} = h\nu$
- 2. Work function
- 3. Stopping potential

De Broglie

The de Broglie wavelength is the wavelength that a particle can have. The wave nature of a particle was proven by diffraction of a crystal. This is because the electron is diffracting which is wave-like.



- 1. $\lambda = \frac{h}{p}$
- 2. de Broglie wavelength
- 3. momentum / mass x velocity

Fluorescent tubes

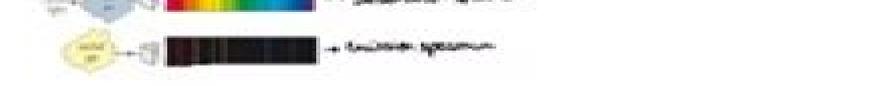


These tubes contain mercury under a high voltage is applied. This excites the gas molecules to which some of the mercury atoms emit photons when they relax. When the gas relaxes it emits photons when the excited mercury atoms emit their energy back into the electron beam, they emit photons in the UV range. A phosphor coating in the tube absorbs the photons and emit visible light.

Energy levels

Electrons can be excited to higher energy levels. Either by absorbing photons from the light or by the collision. In a gas discharge tube, when an electron collides with a gas atom, it can transfer its energy to the atom.

Ionisation energy - is enough energy to pull an electron out of an atom, so will completely change the atom.



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 - The Rise of Digital Reading A Level Physics Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying A Level Physics Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an A Level Physics Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from A Level Physics Notes
 - Personalized Recommendations

- A Level Physics Notes User Reviews and Ratings
- A Level Physics Notes and Bestseller Lists
- 5. Accessing A Level Physics Notes Free and Paid eBooks
 - A Level Physics Notes Public Domain eBooks
 - A Level Physics Notes eBook Subscription Services
 - A Level Physics Notes Budget-Friendly Options
- 6. Navigating A Level Physics Notes eBook Formats
 - ePub, PDF, MOBI, and More
 - A Level Physics Notes Compatibility with Devices
 - A Level Physics Notes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of A Level Physics Notes
 - Highlighting and Note-Taking A Level Physics Notes
 - Interactive Elements A Level Physics Notes
- 8. Staying Engaged with A Level Physics Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers A Level Physics Notes
- 9. Balancing eBooks and Physical Books A Level Physics Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection A Level Physics Notes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine A Level Physics Notes
 - Setting Reading Goals A Level Physics Notes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of A Level Physics Notes
 - Fact-Checking eBook Content of A Level Physics Notes

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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