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web solution 1 the two half reactions balanced as if in acidic solution $2 \text{CN}^- + \text{Au} + \text{CN}^- + 2 \text{e}^- + 2 \text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2 \text{CNO}^-$ make electrons equal convert to basic solution $4 \text{CN}^- + 2 \text{Au} + 2 \text{Au} + \text{CN}^- + 2 \text{e}^-$ multiplied by a factor of 2 $2 \text{e}^- + 2 \text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2 \text{OH}^-$ the final answer

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oxidation number with electrons step 3 balance c e o by adding c e h₂o

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stages first bromine reacts with OH^- ions to form bromide ions and BrO^- ions second the BrO^- ions react to form bromide ions and bromate ions

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