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equation balance the following chemical equations  
 $1 \text{ ch}_4 + \text{o}_2 \rightarrow \text{co}_2 + \text{h}_2\text{o}$   
 $2 \text{ ch}_4 + 2 \text{o}_2 \rightarrow 2 \text{co}_2 + 2 \text{h}_2\text{o}$   
 $2 \text{ na} + \text{cl}_2 \rightarrow 2 \text{nacl}$   
 $3 \text{ al} + \text{o}_2 \rightarrow \text{al}_2\text{o}_3$   
 $4 \text{ al} + 3 \text{o}_2 \rightarrow 2 \text{al}_2\text{o}_3$   
 $4 \text{ n}_2 + 3 \text{h}_2 \rightarrow 2 \text{nh}_3$

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web chemical equations and the law of conservation of matter in the previous section the reaction between hydrogen gas and oxygen gas to produce water in the gaseous phase was shown as a chemical equation  
 $\text{h}_2 + \text{o}_2 \rightarrow \text{h}_2\text{o}$   
at the molecular level the reaction would look something like this notice that there are two oxygen atoms on the

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web nov 9 2023 to balance a chemical equation follow these four simple steps to show the reactants and products write the imbalanced equation calculate the number of atoms of each element on each side of the reaction arrow to make the number of atoms of each element the same on both sides of the equation multiply coefficients the numbers in

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web oct 13 2023 this means that you will need to balance the carbon atoms first 5 use a coefficient to balance the single carbon atom add a coefficient to the single carbon atom on the right of the equation to balance it with the 3 carbon atoms on the left of the equation  
 $\text{c}_3\text{h}_8 + \text{o}_2 \rightarrow \text{h}_2\text{o} + 3 \text{co}_2$