

## Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The charge on an ion is  
a. always positive.  
b. always negative.  
c. either positive or negative.  
d. zero.
- \_\_\_\_\_ 2. According to the octet rule, a calcium atom has a tendency to  
a. lose one electron.  
b. lose two electrons.  
c. gain one electron.  
d. gain two electrons.
- \_\_\_\_\_ 3. If a compound forms by ionic bonding, which is *not* true?  
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.  
b. The net charge of the compound is zero.  
c. The compound contains just two atoms, each of opposite charge.  
d. Several ions group together in a tightly packed structure.
- \_\_\_\_\_ 4. The only property listed that is *not* characteristic of ionic compounds is  
a. high melting point.  
b. hardness.  
c. lack of crystal structure.  
d. good conductor of electricity.
- \_\_\_\_\_ 5. Which formula listed below represents a polyatomic ion?  
a.  $\text{NH}_4^+$   
b.  $\text{H}_2\text{SO}_4$   
c.  $\text{NaCl}$   
d.  $\text{H}_2\text{O}$
- \_\_\_\_\_ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because  
a. ionic substances tend to vaporize at room temperature.  
b. ionic substances are brittle.  
c. attractive forces between ions are greater than the attractive forces between molecules.  
d. the numbers of positive and negative charges are equal in an ionic compound.

# Modern Chemistry Chapter 8 Section 2 Answers

**Frank L. Wiseman**



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## **Table of Contents Modern Chemistry Chapter 8 Section 2 Answers**

1. Understanding the eBook Modern Chemistry Chapter 8 Section 2 Answers
  - The Rise of Digital Reading Modern Chemistry Chapter 8 Section 2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 8 Section 2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Chapter 8 Section 2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter 8 Section 2 Answers
  - Personalized Recommendations
  - Modern Chemistry Chapter 8 Section 2 Answers User Reviews and Ratings
  - Modern Chemistry Chapter 8 Section 2 Answers and Bestseller Lists
5. Accessing Modern Chemistry Chapter 8 Section 2 Answers Free and Paid eBooks
  - Modern Chemistry Chapter 8 Section 2 Answers Public Domain eBooks
  - Modern Chemistry Chapter 8 Section 2 Answers eBook Subscription Services
  - Modern Chemistry Chapter 8 Section 2 Answers Budget-Friendly Options
6. Navigating Modern Chemistry Chapter 8 Section 2 Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Modern Chemistry Chapter 8 Section 2 Answers Compatibility with Devices
  - Modern Chemistry Chapter 8 Section 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 8 Section 2 Answers
    - Highlighting and Note-Taking Modern Chemistry Chapter 8 Section 2 Answers
    - Interactive Elements Modern Chemistry Chapter 8 Section 2 Answers
  8. Staying Engaged with Modern Chemistry Chapter 8 Section 2 Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Modern Chemistry Chapter 8 Section 2 Answers
  9. Balancing eBooks and Physical Books Modern Chemistry Chapter 8 Section 2 Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Modern Chemistry Chapter 8 Section 2 Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Modern Chemistry Chapter 8 Section 2 Answers
    - Setting Reading Goals Modern Chemistry Chapter 8 Section 2 Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Modern Chemistry Chapter 8 Section 2 Answers
    - Fact-Checking eBook Content of Modern Chemistry Chapter 8 Section 2 Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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