

**Section 6.1: Organizing the Elements**

- Group 1 on the periodic table is called **Alkali Metals**
- Group 2 on the periodic table is called **Alkaline Earth Metals**
- Elements in groups 3-12 are called **Transition Metals**
- Group 17 on the periodic table is called **Halogens**
- Group 18 on the periodic table is called **Noble Gases**
- Each period on the periodic table corresponds to a **principal energy level**
- The 2 series that make up the inner transition metals are **Lanthanides** and **Actinides**
- Complete the following chart:

Element	Symbol	Group #	Period #	Group Name	Metal, Metalloid, Nonmetal
Zirconium	<b>Zr</b>	<b>4</b>	<b>5</b>	<b>Transition Metals</b>	<b>Metal</b>
<b>Radium</b>	Ra	<b>2</b>	<b>7</b>	<b>Alkaline Earth Metals</b>	<b>Metal</b>
<b>Bismuth</b>	<b>Bi</b>	15	6	<b>Nitrogen Group</b>	<b>Metal</b>
Gallium	<b>Ga</b>	<b>13</b>	<b>4</b>	<b>Boron Group</b>	<b>Metal</b>

- What are triads? **Sets of 3 elements with similar properties** Whose idea were they? **Dobereiner**
- The modern periodic table is arranged by increasing **atomic number**. Whose idea was this? **Moseley**
- Mendeleev** developed the first periodic table, which was arranged by increasing **atomic mass**
- Determine which of the following pairs will have similar properties (Yes or no):
  - Ca and Be? **Yes – same group (Group 2 – Alkaline Earth Metals)**
  - Sr and Y? **No – not the same group**
  - C and Ge? **Yes – same group (Group 14 – Carbon Group)**

**Section 6.3: Periodic Trends**

- As you go across a period, the atomic radius (size) **decreases** because **the attraction between protons and electrons increases, which makes the atom smaller (opposite of tug of war)**
- As you go down a group, the atomic radius (size) **increases** because **the number of principal energy levels increases**
- Circle the atom with the largest atomic radius and put a square around the atom with the smallest radius:
  - |       |      |    |      |
|-------|------|----|------|
| ○ Li  | □ Ne | F  | B    |
| □ N   | ○ Bi | S  | P    |
| c. Na | ○ Sr | Al | □ Cl |
- Ionization energy is the energy required to **remove** an electron.

# Modern Chemistry Chapter 10 Section 5 Review Answers

**Katrin Zwirgmaier**



## **Modern Chemistry Chapter 10 Section 5 Review Answers:**

*Modern Chemistry* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001      Foundations of Inorganic Chemistry Gary Wulfsberg, 2017-11-02 Foundations of Inorganic Chemistry by Gary Wulfsberg is our newest entry into the field of Inorganic Chemistry textbooks designed uniquely for a one semester stand alone course or to be used in a full year inorganic sequence Foundations of Inorganic Chemistry by Gary Wulfsberg is our newest entry into the field of Inorganic Chemistry textbooks designed uniquely for a one semester stand alone course or to be used in a full year inorganic sequence By covering virtually every topic in the test from the 2016 ACS Exams Institute this book will prepare your students for success The new book combines careful pedagogy clear writing beautifully rendered two color art and solved examples with a broad array of original chapter ending exercises It assumes a background in General Chemistry but reviews key concepts and also assumes enrollment in a Foundations of Organic Chemistry course Symmetry and molecular orbital theory are introduced after the student has developed an understanding of fundamental trends in chemical properties and reactions across the periodic table which allows MO theory to be more broadly applied in subsequent chapters Use of this text is expected to increase student enrollment and build students appreciation of the central role of inorganic chemistry in any allied field Key Features Over 900 end of chapter exercises half answered in the back of the book Over 180 worked examples Optional experiments demos Clearly cited connections to other areas in chemistry and chemical sciences Chapter opening biographical vignettes of noted scientists in Inorganic Chemistry Optional General Chemistry review sections Originally rendered two color illustrations throughout      *Introduction to General, Organic, and Biochemistry* Morris Hein, 1997 The authors recognize that both science and mathematics may be daunting subjects for many students taking this course With this in mind they have anticipated where students might stumble and have paced and organized this text to help them through Their goal is to make the material interesting and relevant so students understand the basic chemical principles related to their career The authors emphasize problem solving and provide a range of practice exercises As in previous editions the text first presents the basic concepts of general chemistry and then moves into organic and biochemistry In this edition the first two sections have been revised primarily to improve explanations and include new pedagogical features The biochemistry portion has been thoroughly updated to include coverage of many recent developments and emerging technologies in the field      *Independent and Weekly Review*, 1914      **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science**, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction      **Energy Research Abstracts**, 1982      **Chemistry** Oxtoby, 1994      The World of Biology P. William

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