

Experiment 10: Electrolysis with a Standard Solution
Introduction: Electrolysis, page 42-43

The purpose of this experiment is to determine the standard potential for various ions. This is done by measuring the standard potential of a half cell against a standard hydrogen electrode. The standard potential of a half cell is defined as the potential of a half cell relative to a standard hydrogen electrode. The standard potential of a half cell is measured by connecting it to a standard hydrogen electrode and measuring the potential difference between the two. The standard potential of a half cell is a measure of the tendency of the half cell to undergo oxidation or reduction. The standard potential of a half cell is a measure of the relative energy of the half cell compared to a standard hydrogen electrode.

In this experiment, you will be determining the standard potential of various half cells and comparing them to the standard potential of a standard hydrogen electrode.

Objectives:

- 1. To determine the standard potential of various half cells.
- 2. To determine the standard potential of the standard hydrogen electrode.
- 3. To determine the standard potential of various half cells.
- 4. To determine the standard potential of various half cells.
- 5. To determine the standard potential of various half cells.

Materials:

- 1. Standard hydrogen electrode
- 2. Standard hydrogen electrode (SHE)
- 3. SHE
- 4. SHE
- 5. SHE
- 6. SHE
- 7. SHE
- 8. SHE
- 9. SHE
- 10. SHE

Procedure:

- 1. Connect the standard hydrogen electrode to the standard hydrogen electrode.
- 2. Measure the standard potential of the standard hydrogen electrode.

Results:

- 1. Standard potential of various half cells.
- 2. Standard potential of various half cells.

Discussion:

- 1. The standard potential of various half cells.
- 2. The standard potential of various half cells.

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