

Name _____ Sex _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then use data for table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount of Excess Reactant
1 molecule	1 molecule	1 molecule	none	none
2 molecules	2 molecules	2 molecules	NO	1 molecule O_2
2 molecules	3 molecules	4 molecules	O_2	1 molecule NO
1.00 mol	2.00 mol	2.00 mol	NO	1.00 mol O_2
2.00 mol	4.00 mol	4.00 mol	NO	2.00 mol O_2
3.00 mol	6.00 mol	6.00 mol	NO	3.00 mol O_2
4.00 mol	8.00 mol	8.00 mol	NO	4.00 mol O_2
5.00 mol	10.00 mol	10.00 mol	NO	5.00 mol O_2
6.00 mol	12.00 mol	12.00 mol	NO	6.00 mol O_2
7.00 mol	14.00 mol	14.00 mol	NO	7.00 mol O_2
8.00 mol	16.00 mol	16.00 mol	NO	8.00 mol O_2
9.00 mol	18.00 mol	18.00 mol	NO	9.00 mol O_2
10.00 mol	20.00 mol	20.00 mol	NO	10.00 mol O_2
11.00 mol	22.00 mol	22.00 mol	NO	11.00 mol O_2
12.00 mol	24.00 mol	24.00 mol	NO	12.00 mol O_2
13.00 mol	26.00 mol	26.00 mol	NO	13.00 mol O_2
14.00 mol	28.00 mol	28.00 mol	NO	14.00 mol O_2
15.00 mol	30.00 mol	30.00 mol	NO	15.00 mol O_2
16.00 mol	32.00 mol	32.00 mol	NO	16.00 mol O_2
17.00 mol	34.00 mol	34.00 mol	NO	17.00 mol O_2
18.00 mol	36.00 mol	36.00 mol	NO	18.00 mol O_2
19.00 mol	38.00 mol	38.00 mol	NO	19.00 mol O_2
20.00 mol	40.00 mol	40.00 mol	NO	20.00 mol O_2

balanced equation mole ratio = 2 mol NO /1 mol O_2

32.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.3125 mol O_2

20.00 g NO = 1 mol NO /30.01 g NO = 0.6664 mol NO

actual mole ratio = 0.6664 mol NO /0.3125 mol O_2 = 2.132 mol NO /1.000 mol O_2

Because the actual mole ratio of NO / O_2 is larger than the balanced equation mole ratio of NO / O_2 , there is an excess of NO / O_2 is the limiting reactant.

Mass of NO used = 0.3125 mol O_2 = 2 mol NO /1 mol O_2 = 0.6250 mol NO

0.6250 mol NO = 30.01 g NO /1 mol NO = 18.75 g NO

Mass of NO_2 produced = 0.6250 mol NO_2 = 46.01 g NO_2 /1 mol NO_2 = 28.75 g NO_2

Excess NO = 20.00 g NO - 18.75 g NO = 1.24 g NO

Name _____ Sex _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium nitride, a substance used in the manufacturing of some semiconductors for lasers.



In one experiment, the reaction yielded 7.42 g of the oxide from a 1.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and GaN are 69.72 g/mol and 122.62 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield volume = 7.42 g GaN
- mass of reactant = 1.00 g Ga
- number of moles of reactant = 1.00 g Ga \times $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ = 7.00 g Ga = 1 mol Ga
- number of moles of product = 0.33 mol Ga \times $\frac{2 \text{ mol GaN}}{1 \text{ mol Ga}}$ = 0.66 mol GaN
- theoretical yield = 0.66 mol GaN \times $\frac{122.62 \text{ g GaN}}{1 \text{ mol GaN}}$ = 80.94 g GaN
- percent yield = $\frac{7.42 \text{ g GaN}}{80.94 \text{ g GaN}} \times 100$ = 9.17%

Study Guide For Content Mastery Answer Key Chapter 12

**Bernadette Mazurek Melnyk, Ellen
Fineout-Overholt**



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Resources in Education, 1975 **Organic Chemistry** Francis A. Carey, 2000 A Market Leading Traditional Approach to Organic Chemistry Throughout all seven editions Organic Chemistry has been designed to meet the needs of the mainstream two-semester undergraduate organic chemistry course This best-selling text gives students a solid understanding of organic chemistry by stressing how fundamental reaction mechanisms function and reactions occur With the addition of handwritten solutions new cutting-edge molecular illustrations updated spectroscopy coverage seamless integration of molecular modeling exercises and state-of-the-art multimedia tools the 7th edition of Organic Chemistry clearly offers the most up-to-date approach to the study of organic chemistry

Instructor Curriculum Guide for the Medical Assistant Mary E Kinn, CPS-CMA-A, Kinn, Woods, 1999-05 Teaching Content to All B. Keith Lenz, Donald D. Deshler, 2004 *Teaching Content to All* includes what every secondary teacher needs to know about instructing students with different learning needs It helps secondary teachers understand academic diversity among students and then plan for and implement instruction that reaches all students The text addresses the unique challenges faced by secondary educators committed to inclusion and to meeting standards for all students *Teaching Content to All* explains research-based teaching techniques and strategies based on understanding instructional goals rather than simply implementing isolated teaching tools Examples are heavily oriented toward the content areas and the planning and teaching routines it presents are easily adaptable across the curriculum by both general and special educators The material can be adapted for the elementary grades

Introduction to Psychology Dennis Coon, 1986 **Study Guide for Obstetrics**

Steven G. Gabbe, Jennifer R. Niebyl, Joe Leigh Simpson, 1992 This book aims to be a convenient easy to use study guide consisting of a series of detailed and probing questions based on each chapter of the book with discussion of the correct and incorrect answers to the questions This edition includes new questions with annotated answers keyed to the larger reference book and should be useful as a preparation tool for obstetrics board examinations Instructor's Resource Guide of the World of Words Margaret Ann Richek, 1989 Psychology Catalog 2005 Neil Thomson, 2004-09 **Psychology** William Buskist, David W. Gerbing, 1990 **Creating Inclusive Classrooms** Spencer J. Salend, 2008 The sixth edition of *Creating Inclusive Classrooms Effective and Reflective Practices for All Students* goes beyond the typical inclusion text in its reflective discussions on how to teach students with varying learning abilities The text contains all of the core information that an inclusion text requires and then takes the reader to a higher level by including issues of gender race ethnicity language socioeconomic status and family structures **Epie Career Education S*e*t*: Analyses of seven hundred prescreened materials**, 1975

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web a tapa spanish pronunciation is an appetizer or snack in spanish cuisine tapas can be combined to make a full meal and can be cold such as mixed olives and cheese or hot such as chopitos which are battered fried baby squid or patatas bravas

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