

# 9

## Thermodynamics

1. Which of the following statements is correct for the spontaneous absorption of a gas? [2014]
- (a)  $\Delta S$  is negative and therefore,  $\Delta H$  should be highly positive  
(b)  $\Delta S$  is negative and therefore,  $\Delta H$  should be highly negative  
(c)  $\Delta S$  is positive and therefore,  $\Delta H$  should be negative  
(d)  $\Delta S$  is positive and therefore,  $\Delta H$  should also be highly positive
2. For the reaction,  $X_2O_4(l) \longrightarrow 2XO_2(g)$ ,  $\Delta U = 2.1 \text{ kcal}$ ,  $\Delta S = 20 \text{ cal K}^{-1}$  at 300 K. Hence,  $\Delta G$  is [2014]
- (a) 2.7 kcal (b) - 2.7 kcal  
(c) 9.3 kcal (d) - 9.3 kcal
3. For a given exothermic reaction,  $K_p$  and  $K_p'$  are the equilibrium constants at temperatures  $T_1$  and  $T_2$ , respectively. Assuming that heat of reaction is constant in temperature range between  $T_1$  and  $T_2$ , it is readily observed that [2014]
- (a)  $K_p > K_p'$   
(b)  $K_p < K_p'$   
(c)  $K_p = K_p'$   
(d)  $K_p = \frac{1}{K_p'}$
4. A reaction having equal energies of activation for forward and reverse reactions has [2013]
- (a)  $\Delta S = 0$  (b)  $\Delta G = 0$   
(c)  $\Delta H = 0$  (d)  $\Delta H = \Delta G = \Delta S = 0$
5. In which of the following reactions, standard reaction entropy changes ( $\Delta S^\circ$ ) is positive and standard Gibbs energy change ( $\Delta G^\circ$ ) decreases sharply with increasing temperature? [2012]
- (a)  $C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow CO(g)$   
(b)  $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$   
(c)  $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$   
(d)  $\frac{1}{2}C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow \frac{1}{2}CO_2(g)$
6. The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is [2012]
- (a) 10.52 cal/mol K (b) 21.04 cal/mol K  
(c) 5.260 cal/mol K (d) 0.526 cal/mol K
7. Standard enthalpy of vaporisation  $\Delta_{\text{vap}}H^\circ$  for water at 100°C is 40.66 kJ mol<sup>-1</sup>. The internal energy of vaporisation of water at 100°C (in kJ mol<sup>-1</sup>) is (assume water vapour to behave like an ideal gas). [2012]
- (a) + 37.56 (b) - 43.76  
(c) + 43.76 (d) + 40.66
8. If the enthalpy change for the transition of liquid water to steam is 30 kJ mol<sup>-1</sup> at 27°C, the entropy change for the process would be [2011]
- (a) 1.0 J mol<sup>-1</sup>K<sup>-1</sup>  
(b) 0.1 J mol<sup>-1</sup>K<sup>-1</sup>  
(c) 100 J mol<sup>-1</sup>K<sup>-1</sup>  
(d) 10 J mol<sup>-1</sup>K<sup>-1</sup>

# Engineering Thermodynamics Important Questions Answers

**Christian G. Meyer**



## **Engineering Thermodynamics Important Questions Answers:**

## **Engineering Thermodynamics Important Questions Answers** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Engineering Thermodynamics Important Questions Answers**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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