

STUDENT BOOK ANSWERS
To come
Chapter 9: Gas laws
Worked example 9.1

- 1 $n = \frac{V}{V_m} = \frac{2.33}{22.71} = 0.1026 \text{ mol}$
 $m = n \times MM = 0.1026 \times (12.01 + (4 \times 1.008)) = 1.646 = 1.65 \text{ g (3sf)}$
- 2 $n = \frac{m}{MM} = \frac{23.0}{14.01 + (3 \times 1.008)} = 1.350 \text{ mol}$
 $V = n \times V_m = 1.350 \times 24.79 = 33.47 = 33.5 \text{ L (3sf)}$

Worked example 9.2

- 1 $2\text{Na(s)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{H}_2\text{(g)}$
 $n(\text{Na}) = \frac{m}{MM} = \frac{1.87}{22.99 + 16.00 + 1.008} = 0.04675 \text{ mol}$
 $n(\text{H}_2) = \frac{1}{2} \times n(\text{Na}) = 0.5 \times 0.04675 = 0.02338 \text{ mol}$
 $V = n \times V_m = 0.02338 \times 24.79 = 0.5795 = 0.580 \text{ L (3sf)}$
- 2 $n(\text{HCl}) = \frac{V}{V_m} = \frac{13.4}{22.71} = 0.5900 \text{ mol}$
 $2\text{KCl(aq)} + \text{H}_2\text{SO}_4\text{(aq)} \rightarrow 2\text{HCl(g)} + \text{K}_2\text{SO}_4\text{(aq)}$
 $n(\text{KCl}) = n(\text{HCl}) = 0.5900 \text{ mol}$
 $m(\text{KCl}) = n \times MM = 0.5900 \times (39.10 + 35.45) = 43.98 = 44.0 \text{ g (3sf)}$

Check your understanding 9.1/9.2

- 1 The volume of a gas changes with temperature and pressure, both of which compress or expand the gas particles.
- 2 Use the formula $n = \frac{V}{V_m}$ to convert the volume of gas to a number of moles. V_m depends upon the temperature of the gas. Use $m = n \times MM$ to convert the number of moles (n) to a mass (m) by multiplying by the molecular mass (MM).
- 3 **a** **i** $2 \times 250 = 500 \text{ mL}$
ii 250 mL
b **i** 500 mL
ii Nothing, as the law of combining volumes only applies to gases.

Chapter 9 Notes Cengage Learning

Marcel A. Müller



Chapter 9 Notes Cengage Learning:

Chapter 9 Notes Cengage Learning Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Chapter 9 Notes Cengage Learning**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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