

CHAPTER 8 REVIEW

Chemical Equations and Reactions

SHORT ANSWER Answer the following questions in the space provided.

2. _____ According to the law of conservation of mass, the total mass of the reacting substances is
- always more than the total mass of the products.
 - always less than the total mass of the products.
 - sometimes more and sometimes less than the total mass of the products.
 - always equal to the total mass of the products.

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 2. In the reaction described by the word equation

$$\text{sodium} + \text{water} \rightarrow \text{sodium hydroxide} + \text{hydrogen gas}$$
 - sodium hydroxide is a product.
 - sodium is a product.
 - water is a reactant.
 - Both (a) and (c).
- _____ 3. Which of the following indicates that a chemical equation is balanced?
 - The numbers of atoms of each element are the same on both sides of the equation.
 - All of the coefficients are the same.
 - The numbers of molecules on each side of the equation are equal.
 - The sums of the coefficients on each side of the equation are equal.
- _____ 4. From a complete and correctly written chemical equation, you can obtain the
 - chemical formulas of the reactants and products.
 - relative amounts of the reactants and products.
 - physical states of the reactants and products.
 - All of the above.
- _____ 5. To balance a chemical equation, you must
 - adjust the subscripts.
 - adjust the number of products.
 - adjust the number of reactants.
 - adjust the coefficients.
- _____ 6. The formation of a solid product in a chemical reaction is represented by the symbol
 - (aq).
 - (g).
 - (l).
 - (s).
- _____ 8. In the expression 3CO_2 , the numbers 3 and 2 are, respectively,
 - a subscript and a coefficient.
 - a coefficient and a subscript.
 - two subscripts.
 - two coefficients.
- _____ 9. What is the total number of molecules that is represented by the following chemical equation? (Add up the coefficients.)

$$\text{NH}_4\text{NO}_2 \rightarrow \text{N}_2 + 2\text{H}_2\text{O}$$
 - 3
 - 4
 - 8
 - 16
- _____ 10. Which equation below violates the law of conservation of mass?
 - $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 - $\text{KCl} + \text{Br} \rightarrow \text{KBr} + \text{Cl}_2$
 - $2\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Fe} + 3\text{CO}_2$
 - $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$
- _____ 11. Which of the following equations represents the balanced equation for the reaction of iron and oxygen?
 - $2\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 - $\text{Fe} + 3\text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 - $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 - $3\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
- _____ 12. Which coefficients correctly balance the formula equation below?

$$\text{KClO}_3(\text{s}) \rightarrow \text{KCl}(\text{s}) + \text{O}_2(\text{g})$$
 - 1, 1, 1
 - 1, 1, 3
 - 2, 2, 3
 - 2, 1, 1
- _____ 13. Which equation is *not* balanced?
 - $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 - $4\text{H}_2 + 2\text{O}_2 \rightarrow 4\text{H}_2\text{O}$
 - $\text{H}_2 + \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$
 - $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- _____ 14. In which type of reaction do two or more compounds react to form one product?
 - synthesis (AKA combination)
 - decomposition
 - combustion
 - single-replacement
- _____ 15. Which of the following is true regarding single-replacement reactions?
 - Only one product is formed.
 - There is only one reagent.
 - One element replaces another in a compound.

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Table of Contents Chapter 8 Review Answers Chemistry

1. Understanding the eBook Chapter 8 Review Answers Chemistry
 - The Rise of Digital Reading Chapter 8 Review Answers Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Review Answers Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 8 Review Answers Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Review Answers Chemistry
 - Personalized Recommendations
 - Chapter 8 Review Answers Chemistry User Reviews and Ratings
 - Chapter 8 Review Answers Chemistry and Bestseller Lists
5. Accessing Chapter 8 Review Answers Chemistry Free and Paid eBooks
 - Chapter 8 Review Answers Chemistry Public Domain eBooks

- Chapter 8 Review Answers Chemistry eBook Subscription Services
- Chapter 8 Review Answers Chemistry Budget-Friendly Options
- 6. Navigating Chapter 8 Review Answers Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 8 Review Answers Chemistry Compatibility with Devices
 - Chapter 8 Review Answers Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 8 Review Answers Chemistry
 - Highlighting and Note-Taking Chapter 8 Review Answers Chemistry
 - Interactive Elements Chapter 8 Review Answers Chemistry
- 8. Staying Engaged with Chapter 8 Review Answers Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 8 Review Answers Chemistry
- 9. Balancing eBooks and Physical Books Chapter 8 Review Answers Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 8 Review Answers Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 8 Review Answers Chemistry
 - Setting Reading Goals Chapter 8 Review Answers Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 8 Review Answers Chemistry
 - Fact-Checking eBook Content of Chapter 8 Review Answers Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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