

## 5.6 Writing Formulas for Ionic Compounds Assignment

### Instructions:

For this assignment, you are being asked to write formulas for Ionic Compounds and for Ionic Compounds containing Polyatomic Ions. Review the examples for each section and then write the formula for the identified compounds.

## Writing Formulas for Ionic Compounds Assignment

### Examples:

Write the formula of potassium bromide.

- $K^{+1}$      $Br^{-1}$
- The sum of the charges is  $+1 + (-1) = 0$ , therefore, the formula is  $KBr$

Write the formula of calcium bromide.

- $Ca^{+2}$      $Br^{-1}$
- The sum of the oxidation numbers is  $+2 + (-1) = +1$ , therefore, we need a second bromide (-1) to obtain a proper formula.
- The new sum is  $+2 + 2(-1) = 0$ , therefore, the formula is  $CaBr_2$ .

A simpler way to understand how to write formulas is to use the crisscross & reduce method. If there is only one atom of an element in the compound, you do not need to put the number 1 behind it. The following are examples:



### Questions 1-6.

Write formulas for the following compounds:

- |                         |         |                         |           |
|-------------------------|---------|-------------------------|-----------|
| 1. sodium and fluorine  | NaF     | 4. aluminum and sulfur  | $Al_2S_3$ |
| 2. magnesium and iodine | $MgI_2$ | 5. hydrogen and oxygen  | $H_2O$    |
| 3. barium and oxygen    | BaO     | 6. lithium and nitrogen | $Li_3N$   |

# Assignment 5 Ionic Compounds

**ML Morrison**



## Assignment 5 Ionic Compounds:

*Chemistry in the Laboratory* James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations *The Analytical Teacher* Victor P. Maiorana, 1994-07

**Using Physical Models of Biomolecules to Teach Concepts of Biochemical Structure in Introductory Undergraduate Chemistry** John Yi, 2004 **Journal of the American Chemical Society** American

Chemical Society, 1960 Proceedings of the Society are included in v 1 59 1879 1937 **Virtual Chemlab** Brian F.

Woodfield, Matthew C. Asplund, Steven Haderlie, 2006 Contains a full virtual lab environment as well as the pre arranged labs that are referenced in the workbook and at the end of the chapter in the textbook Virtual ChemLab can be run directly from the CD or installed on the student's computer *Student Study Guide/solutions Manual to Accompany Foundations of*

*Inorganic, Organic, and Biological Chemistry* Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review

problems Preface *Canadian Journal of Chemistry*, 1961 **Science Abstracts**, 1965 Antimony: Compounds of pentavalent antimony with six, five, and four Sb-C bonds Markus Wieber, 1982 *Holt Chemistry* Ralph Thomas Myers, 2004

*Foundations of College Chemistry* Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years Foundations of College Chemistry Alternate 14th Edition has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They'll learn how to apply concepts with the help of worked out examples In addition

Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis **N Nitrogen** Reinhard Haubold, Peter

Merlet, Claudia Heinrich-Sterzel, Astrid Wietelmann, Ulrike Ohms-Bredeman, Carol Strametz, 1993-10-20 This volume deals with binary nitrogen hydrogen compounds having two three or more nitrogen atoms with the exception of hydrazine and with compounds composed of nitrogen hydrogen and noble gases The important species containing two nitrogen atoms  $N_2H$   $N_2H_2$  and  $N_2H_3$  are described in the first part of this volume Next chains and cycles consisting of three nitrogen atoms are covered Among them hydrogen azide or hydrozoic acid  $HN_3$  is the most extensively studied nitrogen hydrogen compound described in this volume With increasing number of nitrogen atoms the thermochemical stability declines There is however a considerable amount of information on molecules with up to nine linked nitrogen atoms Several of these binary nitrogen

hydrogen compounds could only be isolated in the form of organic derivatives In that case data available for the organic derivatives were included if they were characteristic for the particular unsubstituted N H parent compound

**Rare earth elements: Coordination compounds. sect. 1-4. Coordination compounds** ,1986 *Comptes rendus de l'Académie bulgare des sciences* ,1976

**Introduction to General, Organic, and Biochemistry** Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

*The Elements of Chemistry* Lawrence Powell Eblin, 1965 Chemical concepts and the atomic theory Stoichiometry and the mole concept Atomic numbers and the periodic law Electronic configurations of the elements Chemical bonding and oxidation states Gas laws and the kinetic theory The liquid and solid states Oxygen Hydrogen and water Acids and bases Solutions and colloids Reaction rates and equilibrium Oxidation reduction and electrochemistry Energy and chemical change Organic compounds Biochemical compounds Transmutation of the elements

Australian Journal of Chemistry ,1985 *Russian Journal of Inorganic Chemistry* ,1990 **Abstracts of Papers** American Chemical Society. Meeting,

**General Chemistry** Ralph Sherman Becker, Wayne Ernest Wentworth, 1980

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