

<b>Tax Rate</b>	<b>Single Filers</b>	<b>Joint Filers</b>	<b>Head of Household Filers</b>
10%	Under \$11,000	Under \$22,000	Under \$15,700
12%	Over \$11,000	Over \$22,000	Over \$15,700
22%	Over \$44,725	Over \$89,450	Over \$59,850
24%	Over \$95,375	Over \$190,750	Over \$95,350
32%	Over \$182,100	Over \$364,200	Over \$182,100
35%	Over \$231,250	Over \$462,500	Over \$231,250
37%	Over \$578,125	Over \$693,750	Over \$578,100

# Tax Bracket Review

**R Pring**



## Tax Bracket Review:

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction

of 8.001 grams of copper(II) sulfate ( $\text{CuSO}_4$ ) and 2.0153 grams of iron powder ( $\text{Fe}$ ) the goal of this experiment was to determine the product of copper(II) sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8.00 grams of copper(II) sulfate onto a piece of filter paper taking the filter paper's mass into account add these to the beaker 3 measure 50.0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper(II) sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper(II) sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper(II) sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper(II) sulfate in the iron copper(II) sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper(II) sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper(II) sulfate ions the copper(II) is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated  $\text{CuSO}_4$  solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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web stoichiometry the reaction of iron with copper ii sulfate purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction

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