

# Calculations need to be based on the limiting reactant.

- Example 2: What is the maximum mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?



$$\frac{95.6 \text{ g } CS_2}{76.2 \text{ g } CS_2} \times \frac{1 \text{ mol } CS_2}{1 \text{ mol } CS_2} \times \frac{2 \text{ mol } SO_2}{1 \text{ mol } CS_2} \times \frac{64.1 \text{ g } SO_2}{1 \text{ mol } SO_2} = 161 \text{ g } SO_2$$

$$\frac{100. \text{ g } O_2}{32.0 \text{ g } O_2} \times \frac{1 \text{ mol } O_2}{3 \text{ mol } O_2} \times \frac{2 \text{ mol } SO_2}{1 \text{ mol } O_2} \times \frac{64.1 \text{ g } SO_2}{1 \text{ mol } SO_2} = 134 \text{ g } SO_2$$

Which reactant is **LIMITING**?

# Limiting Reactant Problems And Solutions

**Clifford Lane**



## **Limiting Reactant Problems And Solutions:**

Uncover the mysteries within Explore with is enigmatic creation, Discover the Intrigue in **Limiting Reactant Problems And Solutions** . This downloadable ebook, shrouded in suspense, is available in a PDF format ( Download in PDF: \*). Dive into a world of uncertainty and anticipation. Download now to unravel the secrets hidden within the pages.

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