

NAME _____ DATE _____ CLASS _____

CHAPTER 12 **STUDY GUIDE FOR CONTENT MASTERY**

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , N_2 , and NO are 32.00 g/mol, 28.02 g/mol, and 30.01 g/mol, respectively.

Amount of O_2	Amount of N_2	Amount of NO	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	N_2	2 molecules O_2
2 molecules	4 molecules	4 molecules	4 O_2	4 molecules NO
1.00 mol	2.00 mol	2.00 mol	none	none
4.00 mol	8.00 mol	8.00 mol	N_2	2.00 mol O_2
0.05 mol	0.10 mol	0.10 mol	N_2	1.95 mol O_2
1.00 mol	2.00 mol	2.00 mol	O_2	2.00 mol NO
0.200 mol	0.200 mol	0.400 mol	N_2	0.400 mol O_2
32.00 g	56.04 g	92.02 g	none	none
64.00 g	112.08 g	184.04 g	O_2	32.00 g NO
64.00 g	56.04 g	92.02 g	O_2	1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

16.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.5125 mol O_2

28.02 g N_2 = 1 mol N_2 /28.02 g N_2 = 0.9994 mol N_2

actual mole ratio = 0.9994 mol N_2 /0.5125 mol O_2 = 2.132 mol N_2 /1.000 mol O_2

Because the actual mole ratio of N_2 / O_2 is larger than the balanced equation mole ratio of N_2 / O_2 , there is an excess of N_2 ; O_2 is the limiting reactant.

Mass of NO used = 0.5125 mol O_2 = 2 mol NO /1 mol O_2 = 0.4250 mol NO

0.4250 mol NO = 30.01 g NO /1 mol NO = 12.76 g NO

Mass of NO_2 produced = 0.4250 mol NO_2 = 46.01 g NO_2 /1 mol NO_2 = 20.76 g NO_2

Excess N_2 = 28.02 g N_2 = 18.76 g N_2 = 1.24 g N_2

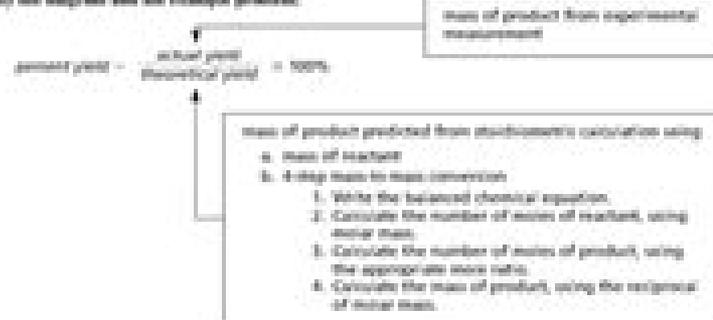
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CHAPTER 12 **STUDY GUIDE FOR CONTENT MASTERY**

Section 12.4 Percent Yield

In your textbook, read about the percent of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of yellow oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: 7.00 g Ga = $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ 7.00 g Ga = 1 mol Ga/69.72 g Ga
- number of moles of product: 0.100 mol Ga = $\frac{2 \text{ mol Ga}_2\text{O}_3}{1 \text{ mol Ga}}$ 0.100 mol Ga = 2 mol Ga_2O_3 /4 mol Ga
- theoretical yield: 0.200 mol Ga_2O_3 = $\frac{187.44 \text{ g Ga}_2\text{O}_3}{1 \text{ mol Ga}_2\text{O}_3}$ correct
- percent yield: $\frac{6.97 \text{ g Ga}_2\text{O}_3}{7.42 \text{ g Ga}_2\text{O}_3} \times 100$ 7.42 g Ga_2O_3 /6.97 g Ga_2O_3 = 100

Standardized Test Practice Chapter 12 Chemistry Answers

WJ Hussar



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