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web the electron removed during the ionization of beryllium $He\ 2s^2$ is an s electron whereas the electron removed during the ionization of boron $He\ 2s^2\ 2p^1$ is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^1$ or simply put $Ar\ 4s^1$ q2 what is the total number of electrons that can be filled in s p and d subshells

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