

## 5

## ELECTRONS IN ATOMS

**A. Matching**

Match each term in Column B with the correct description in Column A. Write the letter of the correct term on the line.

Column A	Column B
_____ 1. the ways in which electrons are arranged around the nuclei of atoms	a. Aufbau principle
_____ 2. the ejection of electrons by metals when light strikes on them	b. quantum
_____ 3. the region around the nucleus of an atom where an electron is likely to be locating	c. photoelectric effect
_____ 4. An atomic orbital may describe at most two electrons.	d. atomic orbitals
_____ 5. the regions within which electrons have the highest probability of being found	e. Pauli exclusion principle
_____ 6. When electrons occupy orbitals of equal energy, one electron enters each orbital and all the orbitals contain one electron.	f. energy level
_____ 7. the amount of energy required to move an electron from its present energy level to the next highest one	g. Hund's rule
_____ 8. light quanta	h. electron configurations
_____ 9. Electrons enter orbitals of lowest energy first.	i. quantum mechanical model
_____ 10. the modern description of the electrons in atoms	j. photons

**B. Multiple Choice**

Choose the best answer that applies to letter on the list.

- \_\_\_\_\_ 11. Bohr's contribution to the development of atomic structure:
- was referred to as the "planetary model."
  - was the discovery that electrons surrounded a dense nucleus.
  - was proposed that electrons travel in circular orbits around the nucleus.
  - is the quantum mechanical model.
- \_\_\_\_\_ 12. What is the total number of orbitals in the third principal energy level?
- 1
  - 2
  - 4
  - 16

# Electrons In Atoms Chapter 5 Answer Key

**Minjie Lin**



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