

Chapter 10 Guided Notes

Name _____

1. _____ attractions are bonds that hold compounds together
 - a. What are 2 examples of these attractions?
2. _____ attractions are between molecules.
3. Which attraction is stronger?
4. Which state of matter will have the **strongest** intermolecular forces? Which state has the **weakest**?



5. Order the intermolecular forces (IMF) in order of weakest to strongest: hydrogen bonding, dipole-dipole, ion-dipole, London dispersion
6. List all the possible IMF for the following compounds:
 - a. HCl
 - b. NH₃
 - c. H₂O
 - d. CH₄
 - e. NaCl dissolved in water

7. For each type of compound determine what forces are possible:

Type of IMF	Atoms	Nongpolar molecules	Polar molecules without OH, NH, or HF groups	Polar molecules containing OH, NH, or HF groups	Ionic solids dissolved in polar liquids
Dispersion forces					
Dipole-dipole					
Hydrogen bonding					
Ion-dipole					

Explore the following procedure and complete the subsequent practice exercises in preparation for the synchronous class discussion. Step 1 or Step 2 are review from general chemistry I. Review Lewis structures

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