

Exam-style questions and sample answers have been written by the author. In examinations, the way marks are awarded may be different.

# Teacher's resource answers

## Chapter 15

1 a Isomers are compounds with the same molecular formula [1]

but different structural formulae. [1]

b [2]



2 a heat change = mass  $\times$  specific heat capacity  $\times$  change in temperature [1]  
 $= 200 \times 4.18 \times 30$  [1]  
 $= 25\,080 \text{ J} = 25.1 \text{ kJ}$  [1]

b number of moles of 2-methylbutane [1]  
 $= \frac{0.626}{72} = 8.69 \times 10^{-3} \text{ mol}$  [1]

standard enthalpy of combustion [1]  
 $= \frac{25.1}{8.69 \times 10^{-3}}$  [1]

$= -2890 \text{ kJ mol}^{-1}$  [1]

c heat change is the same, 25.1 kJ [1]  
 number of moles of 2,2-dimethylpropane

$= \frac{171}{24000} = 7.13 \times 10^{-3}$  [1]

standard enthalpy of combustion [1]  
 $= \frac{25.1}{7.13 \times 10^{-3}}$  [1]

$= -3520 \text{ kJ mol}^{-1}$  [1]



cycle with carbon dioxide and water as the third corner and where the directions of the three arrows are correct, and [1]

labelling arrows with correct enthalpies [1]

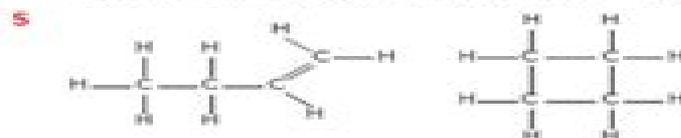
$\Delta H_{\text{combustion}} + (-3520) = -2890$  [1]

$\Delta H_{\text{combustion}} = -2890 - (-3520) = +630 \text{ kJ mol}^{-1}$  [1]

4 The 2-methylbutane has a greater surface area. [1]

therefore a greater number of van der Waals' forces/induced dipoles. [1]

It has stronger intermolecular forces and therefore more energy required to separate molecules (and a higher boiling point). [1]



1 mark for each structure

Note: Carbon and hydrogen atoms are shown for displayed formulae. [2]

6 a  $\text{C}_2\text{H}_6 + \text{Br}_2 \rightarrow \text{C}_2\text{H}_5\text{Br} + \text{HBr}$  [1]  
 correct reactants [1]

for correct products [1]

b Initiation. [1]

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