

GRAVIMETRIC ANALYSIS PROBLEMS - EXERCISES IN STOICHIOMETRY

1. In the analysis of 0.7011 g of an impure chloride containing sample, 0.9805 g of AgCl were precipitated. What is the percentage by mass chloride in the sample?

$$\begin{aligned}m_{\text{sample}} &= 0.7011 \text{ g} \\m_{\text{AgCl}} &= 0.9805 \text{ g} \\m_{\text{Cl}} &= 0.9805 \text{ g} \times \frac{35.453 \text{ g/mol}}{143.32 \text{ g/mol}} \\&= 0.2425 \text{ g Cl} \\ \% \text{Cl} &= \frac{0.2425 \text{ g Cl}}{0.7011 \text{ g}} \times 100 = 34.59\% \text{Cl}\end{aligned}$$

2. A 0.4054 g solid organic sample containing covalently bound bromide and no other halogens was placed in a porcelain crucible with about one gram of fresh sodium metal. In a process known as *sodium fusion*, the mixture was heated in a furnace to 450°C which charred and vaporized the organic portion of the molecule and converted the covalent bromide into sodium bromide. Excess sodium was decomposed by adding small portions of water which also dissolved the sodium bromide. The clear solution was quantitatively transferred to a beaker, acidified with dilute nitric acid, and diluted to about 50 mL. A 6 mL quantity of 0.1 M AgNO₃ was added to the solution and the mixture heated to about 60°C for an hour to age and digest the precipitate. After filtering, the mass of the silver bromide produced was determined to be 37.8 mg. What is the percentage by mass bromine in the organic compound?

$$\begin{aligned}m_{\text{sample}} &= 0.4054 \text{ g} = 405.4 \text{ mg} \\m_{\text{AgBr}} &= 37.8 \text{ mg} \\ \% \text{Br} &= \frac{37.8 \text{ mg} \times \frac{79.904 \text{ g/mol}}{187.77 \text{ g/mol}}}{405.4 \text{ mg}} \times 100 \\&= 3.97\% \text{ Br}\end{aligned}$$

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