

## General Chemistry Questions

### Electronic Structure and Periodic Table

1. What value or values of  $m_l$  are allowable for an orbital with  $l = 2$ ?
  - a. 0
  - b. 2
  - c. -1
  - d. none of the above
  - e. all of the above
2. According to Bohr Theory, which of the following transitions in the hydrogen atom will give rise to the **least energetic** photon?  
Use the equation:  $E_n = (-2.18 \times 10^{-18} \text{ J})(1/n^2)$ 
  - a.  $n = 5$  to  $n = 3$
  - b.  $n = 6$  to  $n = 1$
  - c.  $n = 4$  to  $n = 3$
  - d.  $n = 5$  to  $n = 4$
  - e.  $n = 6$  to  $n = 5$
3. Consider a  $3d_{zz}$  orbital. Which of the following statements is **incorrect**?
  - a. The  $xz$  plane is a nodal surface.
  - b. The  $xz$  plane divides the electron probability distribution into two identical mirror-image halves.
  - c. The  $xy$  plane divides the electron probability distribution into two identical mirror-image halves.
  - d. The  $yz$  plane divides the electron probability distribution into two identical mirror-image halves.
  - e. The nucleus is located at a node.
4. The electronic configuration of the element whose atomic number is 26 is:
  - a.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^0 3d^6$
  - b.  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$
  - c.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
  - d.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6 4p^2$
  - e. none of the above
5. Which of the following has the largest radius?
  - a. F
  - b. N
  - c. C
  - d. O
  - e. Ne

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