

Ideal gas law - problems and solutions

1. Ideal gases in a closed container initially have volume V and temperature T . The final temperature is $5/4T$ and the final pressure is $2P$. What is the final volume of gases.

Known :

Initial volume (V_1) = V

Initial temperature (T_1) = T

Final temperature (T_2) = $5/4 T$

Initial pressure (P_1) = P

Final pressure (P_2) = $2P$

Wanted : Final volume (V_2)

Solution :

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\frac{P V}{T} = \frac{(2P) V_2}{5/4 T}$$

$$\frac{V}{1} = \frac{2 V_2}{5/4}$$

$$5/4 V = 2 V_2$$

$$V_2 = \frac{5V}{4} \cdot \frac{2}{1}$$

$$V_2 = \frac{5V}{4} \times \frac{1}{2}$$

$$V_2 = \frac{5V}{8}$$

2. Determine the volume of 2.00 moles of gases (ideal gas) at STP. STP = Standard Temperature and Pressure.

Known :

Moles of gas (n) = 2 moles

Standard temperature (T) = $0^\circ\text{C} = 0 + 273 = 273$ Kelvin

Standard pressure (P) = 1 atm = 1.013×10^5 Pa

Universal gas constant (R) = 8.315 Joule/mole.Kelvin

Wanted : Volume of gases (V)

Solution :

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<https://physics.gurumuda.net/ideal-gas-law-problems-and-solutions.htm>

Gas Law Problems With Solutions

**Robert J. Salm, Robert J. Salm Robert J.
Salm Ph.D.**

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Dr. Salm's Chemistry Problem Drill Book Robert J. Salm, Robert J. Salm Robert J. Salm Ph.D., 2005 This drill book contains many common problem types that are asked in General Chemistry classes in High School and College This work will give you practice with the major problem types as you prepare for finals and standardized tests **Survival Guide to General Chemistry** Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields *Journal of Chemical Education*, 1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers *Math Concepts for Food Engineering* Richard W. Hartel, D.B. Hyslop, Robin K. Connelly, T.A. Howell Jr., 2008-03-17 A Supplement for Food Science Engineering Students Who Need to Improve Their Mathematical Skills A remedial textbook for understanding mathematical theories and formulas *Math Concepts for Food Engineering Second Edition* helps students improve their mathematical skills so that they can succeed in food engineering cour **Physico-chemical Calculations** Joseph Knox, 1919 **Questions and Answers on Federal Tax Laws, Problems Affecting Individuals, Partnerships and Corporations, January, 1924** Irving Bank-Columbia Trust Company, 1924 University High School Journal, 1925 *Student Solutions Manual for Chang's Chemistry* Philip C. Keller, Raymond Chang, Jill L. Keller, 1988 *Ebook: Chemistry: The Molecular Nature of Matter and Change* Silberberg, 2015-01-16 *Ebook Chemistry The Molecular Nature of Matter and Change* Physico-chemical Calculations Joseph Knox, 1944 **Chemical Symbolism and Calculations** Stanley Wallace Morse, 1953 **Chemical Calculations** Raymond Harman Ashley, 1929 *Journal* University High School (Oakland, Calif.), 1926 *The Journal of Physical Chemistry*, 1912 The Journal of Physical Chemistry Wilder Dwight Bancroft, 1912 Includes section New Books **Nature** Sir Norman Lockyer, 1917 **Introduction to the Study of Chemical Philosophy** Sir William Augustus

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