

# Buffers Keep the Balance

## Post-Lab Questions

1. Compare the measured pH value for the model carbonate blood buffer to
  - a. the expected pH of an ideal carbonic acid-bicarbonate buffer, and
  - b. the actual pH of the carbonic acid-bicarbonate buffer system present in the blood.
2. Based on the pH comparisons in Question #1, which solution, the model carbonate blood buffer or an actual blood buffer, is more likely to contain a greater proportion of the carbonic acid component of the buffer compared to the bicarbonate component? Explain your reasoning.
3. What are the effects of adding even a small amount of HCl or NaOH on the pH value of the control solution (water)? Compare this to the effect of adding HCl or NaOH to the model carbonate blood buffer.
4. Which phosphate buffer solution in Part B corresponds to the composition of an ideal buffer solution? Compare its measured pH value with the calculated pH of the ideal buffer solution.

# Buffers Keep The Balance Lab Answers

**G Psacharopoulos**



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