

Worksheet 1-1 - Measuring Reaction Rates

Name - _____
Due Date - _____

Answer the questions in the space provided. You must show all of your work to receive full marks. All answers MUST be rounded to the correct number of Sig. Figures.

1. A chemist wishes to determine the rate of reaction of zinc with hydrochloric acid. The equation for the reaction is:



A piece of zinc is dropped into 1.00 L of 0.100 M HCl and the following data were obtained:

Time	Mass of Zinc
0 s	0.016 g
4 s	0.014 g
8 s	0.012 g
12 s	0.010 g
16 s	0.008 g
20 s	0.006 g

$$\begin{array}{r} 0.016\text{g} \\ - 0.006\text{g} \\ \hline 0.010\text{g} \end{array}$$

- a) Calculate the *Rate of Reaction* in grams of Zn consumed per second.

$$\text{rate} = \frac{\Delta g}{\Delta \text{time}} = \frac{0.010\text{g}}{20\text{sec}}$$

Answer 5.0×10^{-4}

- b) Calculate the *Rate of Reaction* in moles of Zn consumed per second.

$$5.0 \times 10^{-4} \frac{\text{g}}{\text{sec}} \times \frac{1\text{mol}}{65.4\text{g}} =$$

Answer $7.6 \times 10^{-6} \frac{\text{mol}}{\text{sec}}$

- c) Write out the complete ionic equation for the reaction.



- d) What will happen to the $[\text{H}^+]$ as the reaction proceeds?

decrease

- e) What will happen to the $[\text{Cl}^-]$ as the reaction proceeds?

nothing, it is a spectator

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Ralph J. Mill



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Chemistry Rob Ritchie, 2004 These New editions of the successful highly illustrated study revision guides have been fully updated to meet the latest specification changes Written by experienced examiners they contain in depth coverage of the key information plus hints tips and guidance about how to achieve top grades in the A2 exams Progress check questions test recall and understanding and end of unit sample questions and model answers provide essential practice to improve students exam technique **McGraw-Hill's 500 MCAT General Chemistry Questions to Know by Test Day** John T.

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