

# 13.2 THE NATURE OF LIQUIDS

## Section Review

### Objectives

- Identify factors that determine physical properties of a liquid
- Define *evaporation* in terms of kinetic energy
- Describe the equilibrium between a liquid and its vapor
- Identify the conditions under which boiling occurs

### Vocabulary

- vaporization
- evaporation
- vapor pressure
- boiling point
- normal boiling point

### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Liquids are much 1 than gases. Liquids and solids are known as 2 states of matter. The conversion of a liquid to a gas or vapor is called 3. When a liquid that is not 4 changes to a gas, the process is called evaporation. A liquid evaporates faster when heated; however, evaporation itself is a 5 process. When a partially filled container of liquid is sealed, some of the particles at the 6 of the liquid vaporize. These particles collide with the walls of the container, producing a force called 7. The vapor pressure of a liquid can be determined by a device called a 8. A liquid boils when its 9 equals the external pressure. The normal boiling point of a liquid is the temperature at which the vapor pressure is equal to 10.

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### Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- \_\_\_\_\_ 11. Particles in a liquid don't have enough kinetic energy to overcome the attractive forces between them and vaporize.

# Nature Of Liquids Section Review Key

**Terry C. Jones**



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