



Understanding Business Law

9TH EDITION

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Business Law 9th Edition

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Business Law: (As per latest syllabuses of universities in India), 9th Edition Kuchhal M.C. & Kuchhal Vivek, This book deals with the fundamental branches of business law namely Law of Contract Law of Sale of Goods Law of Partnership Law of Negotiable Instruments and Law of Limited Liability Partnership LLP Business Law is intended to serve as a textbook for the students of BCom BCom Hons CA Foundation Course CA Intermediate Course CS Foundation Programme ICMA Intermediate BBA MBA and also for those appearing for banking and competitive examinations

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$$l \cdot m \left(\frac{1000 \cdot m}{1000 \cdot d} \right)$$

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web molarity has units of $\frac{\text{mol}}{\text{L}}$ which can be abbreviated as molar or M pronounced molar the molar concentration of the solute is sometimes abbreviated by putting square brackets around the chemical formula of

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web molarity is the number of moles of a substance per litre of solution also known as molar concentration a capital M signifies solutions labelled with molar concentration a 1.0 M solution contains 1 mole of solute per litre of solution molality is the number of solute moles per kilogram

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web 1. molarity $M = \frac{m}{V}$ mass molar mass $\times 0.4500 \text{ L} = 0.825 \text{ g} \times \frac{141.9579 \text{ g/mol}}{0.0129 \text{ m}^3} = 0.825 \text{ g} \times \frac{141.9579 \text{ g/mol}}{0.00581158 \text{ m}^3} = 0.825 \text{ g} \times \frac{141.9579 \text{ g/mol}}{0.00581158 \text{ m}^3} = 20.3 \text{ M}$
 2. mole fraction $n_2 = 0.825 \text{ g} \times \frac{141.9579 \text{ g/mol}}{0.00581158 \text{ m}^3} = 20.3 \text{ M}$
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web apr 7 2014 molarity moles of solute liters of solution $8.4 \times 10^{-2} \text{ mol} \times \frac{1000 \text{ mL}}{250 \text{ mL}} = 0.336 \text{ mol/L}$
 calculate molarity 5 moles / 0.25 liters = 20 M
 a solution with molarity 2 requires 2 M of NaOH per liter so $4 \times 28 \text{ g} = 112 \text{ g}$
 a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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web answer the correct option is C explanation the molarity of a mixture can be calculated using the following formula
 $M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_1 + V_2 + \dots + V_n}$
 $1.2 \times 480 + 1.5 \times 520 = 1.344 \text{ M}$
 with increase in temperature which of these changes a molality b fraction of solute present in water

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web 1. calculate the molarity of a solution which contains 0.40 mol of a substance dissolved in 1.6 L of a solution
 2. what is the molarity of a solution containing 325 g of NaCl dissolved in 750 mL of solution
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 4. 724 g of ammonium phosphate in 4500 mL of alcohol

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web determine the molarity for each of the following solutions
 0.444 mol of CO_2 in 0.654 L of solution
 98.0 g of phosphoric acid H_3PO_4 in 1.00 L of solution
 0.2074 g of calcium hydroxide Ca(OH)_2 in 40.00 mL of solution
 10.5 kg

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