

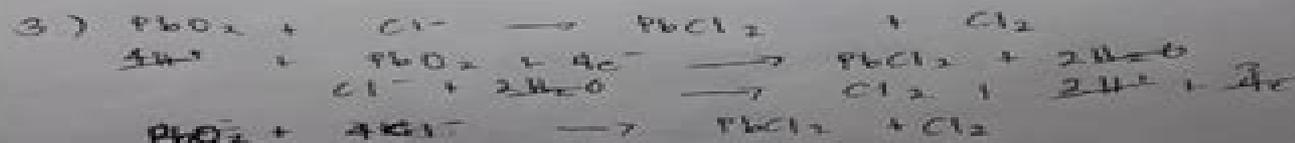
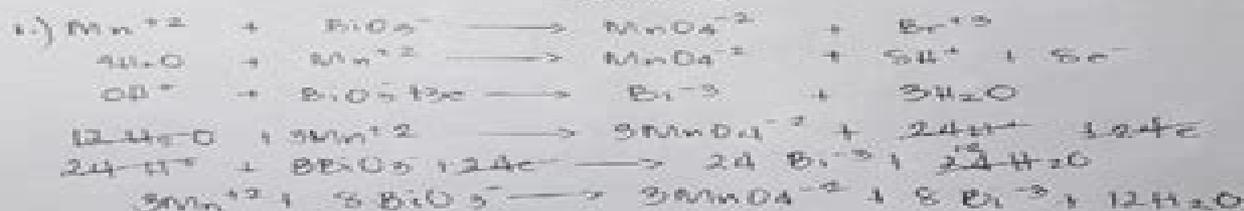
## OXIDATION - REDUCTION REACTION

A. Balance the following ionic equations occurring in acidic solution.

1.  $\text{Mn}^{+2} + \text{BiO}_3^- \rightarrow \text{MnO}_4^{+2} + \text{Bi}^{+3}$
2.  $\text{NiO}_2 + \text{Mn}(\text{OH})_2 \rightarrow \text{Mn}_2\text{O}_3 + \text{Ni}(\text{OH})_2$
3.  $\text{PbO}_2 + \text{Cl}^- \rightarrow \text{PbCl}_2 + \text{Cl}_2$
4.  $\text{H}_2\text{PO}_3 + \text{Cr}_2\text{O}_7^{+2} \rightarrow \text{H}_2\text{PO}_4 + \text{Cr}^{+3}$
5.  $\text{AsO}_3^{+2} + \text{Br}_2 \rightarrow \text{AsO}_4^{+2} + \text{Br}^-$

### Oxidation - Reduction Reaction

A. Balance the following ionic equations occurring in acidic solution



# Oxidation And Reduction Practice Problems Answers

**George Gilbert Long, Forrest C. Hentz**



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