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**Paul McFedries**



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choose 1 answer the solution becomes negatively charged due to the majority  $Cl^-$  ions a the solution becomes negatively charged due to the majority  $Cl^-$  ions the solution becomes positively charged due to the stronger  $Mg^{2+}$  ions b the solution becomes positively charged due to

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choose the correct answer from the options below the spheres will swing towards each other due to the attractive electrostatic force between them

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question titlesolution answer d somewhere in region d justification with the test charge and c 1 being negative there is a repulsive force on the test charge to the right from c

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session 9 electrostatics key concepts in this session we will focus on summarising what you need to know about electrostatics and types of charges electric fields properties

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before solving a question first of all look at what quantities are given and then think of a strategy to solve the question books for electrostatics first of all read the ncert book each and

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1 depends on the magnitude of charges 2 force between two charges may be attractive or repulsive 3 it depends on the medium 4 applicable only for charges at rest gravitational

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outside express your answer in terms of the total charge  $q$  on the surface  $p = z \cos \theta$   $r \sin \theta = z \cos \theta$   $q$  figure 2.4 problem 2.7 consider a slice of the shell centered on the  $z$  axis see

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formula for molality molality = moles of solute / mass of solvent in kg example of molality

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web correct answer c both a and b q2 the density of 1 M solution of HCl is 1.0585 g/ml the molality of the solution is 1.0585 1.010 0.0585 correct answer b 1 explanation molality 
$$m = \frac{M \times 1000}{1000d}$$

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web sep 13 2023 the only difference between molality and molarity is in the denominator molality is measured in kilograms of solvent whereas molarity is measured in liters of solution when studying the properties of solutions related to vapour pressure and temperature changes concentrations expressed in molality are used

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web answer the correct option is c explanation the molarity of a mixture mmix can be calculated using the following formula  $m_{mix} = \frac{m_1 v_1 + m_2 v_2}{v_1 + v_2}$  1 5 x 480 1 2 x 520 480 520 720 624 1000 1 344 m q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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