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Business Overview

Weekly Sales

02/10/2008-02/16/2008	3012.63 USD
02/03/2008-02/09/2008	4255.34 USD
01/27/2008-02/02/2008	3489.47 USD

Jan 2008 Overview

Ending Balance	17,504.38 USD
Payments Received	15,657.64 USD

New Reporting Tools

Save yourself time.

- Track sales trends easily
- Reconcile your accounts faster
- Search transactions quickly

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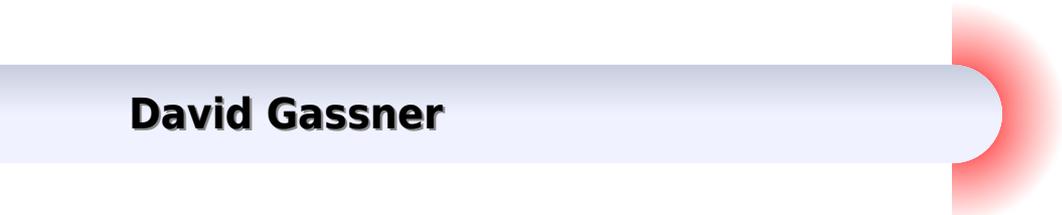
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David Gassner



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readers through a wide ranging array of features Note this isn't a staid guide to generating traditional reports but rather it shows you how the Ruby language and Rails framework can create truly compelling reporting services by plugging into popular third party applications and services such as Google AdWords UPS.com iTunes and Salesforce.com

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h₂o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + \text{Fe}^{2+} + \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + 6\text{Fe}^{2+} + 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web balance each half reaction in basic solution 4 5 6 7 $\text{Cr}_2\text{O}_7^{2-} + \text{NO}_2 + \text{SO}_4^{2-} + \text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method 8 $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ + \text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{Te} + 10\text{H}_2\text{O} + 11\text{PbO}_2 + \text{IO}_2 + \text{I}_2 + \text{Pb} + 12\text{H}^+$ as $\text{NO}_3^- + \text{IO}_3^- + \text{Re} + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H_2O H^+ and e^- in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 + \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b $\text{C}_6\text{H}_8\text{O}_6$ and $\text{C}_6\text{H}_6\text{O}_6$ c H_2O and $\text{C}_6\text{H}_8\text{O}_6$ d $\text{C}_6\text{H}_8\text{O}_6$ and $\text{C}_6\text{H}_6\text{O}_6$