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Marco Cascella



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differences and similarities with hand warmers leading to research on how they work children can investigate why materials

have been chosen for the hand warmers linking to previous knowledge on materials and properties

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web $m \times c \times \Delta t$ equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution

and Δt is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of

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web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves

exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH}(s) \rightarrow \text{NaOH}(aq)$ $\Delta H_{\text{soln}} = 44.2 \text{ kJ}$

mol rxn write equations to similarly represent the dissolving process for

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web hand warmers are usually placed inside the gloves to keep your hands warm the heat keeping your hands warm is

formed as a result of an exothermic reaction which occurs in a small packet hand warmers have 5 main ingredients in them

salt carbon polypropylene vermiculite water and iron

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web lab investigation in today s experiment you will design and execute an experimental procedure to determine which of three ionic compounds is most suitable for use in a hand warmer the three compounds you will use is determined by the group number you are assigned the ideal hand warmer