

experiment date(s):

Experiment 6 - Analysis of an Unknown Chloride

Goal: Determine the % chloride (Cl) in an unknown sample.

In the experiment, we will titrate with silver nitrate (AgNO_3) that has a known concentration.

net ionic equation: $\text{Ag}^+_{(aq)} + \text{Cl}^-_{(aq)} \longrightarrow \text{AgCl}_{(s)}$ white precipitate

When Cl^- runs out, the reaction is complete; this is called the **end point**. We will know the end point is reached because CrO_4^{2-} (chromate) is added to the chloride solution. It will also form a precipitate with Ag^+ , but only after all the Cl^- is used up

net ionic equation: $2 \text{Ag}^+_{(aq)} + \text{CrO}_4^{2-}_{(aq)} \longrightarrow \text{Ag}_2\text{CrO}_4_{(s)}$ red precipitate

The chromate ion, yellow in color, will make the titration appear to form yellow precipitate. We will stop adding the silver solution when the chloride solution begins to turn red.

ex. A solid chloride sample weighing 0.3361 g requires 61.46 mL of 0.06258 M AgNO_3 to reach the end point.

(a) How many mol of Cl^- are present in the sample?

$$\frac{0.06146 \text{ L AgNO}_3}{1 \text{ L AgNO}_3} \left(\frac{0.06258 \text{ mol AgNO}_3}{1 \text{ L AgNO}_3} \right) \left(\frac{1 \text{ mol Cl}^-}{1 \text{ mol AgNO}_3} \right) = 3.846 \times 10^{-3} \text{ mol Cl}^-$$

(b) How many g of Cl^- are present in the sample?

$$\frac{3.846 \times 10^{-3} \text{ mol Cl}^-}{1 \text{ mol Cl}^-} \left(\frac{35.45 \text{ g Cl}}{1 \text{ mol Cl}} \right) = 0.1363 \text{ g Cl}^-$$

(c) What is the mass % of Cl^- in the sample?

$$\frac{\text{g Cl}^-}{\text{g sample}} \times 100 = \frac{0.1363 \text{ g}}{0.3361 \text{ g}} \times 100 = 40.55 \%$$

Procedure: omit back titration with NaCl

Waste disposal: Dispose of all liquids and rinse all solids into the Heavy Metal Inorganic waste container.

ASA: all

Analysis Of An Unknown Chloride Answers

D Kirk



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Table of Contents Analysis Of An Unknown Chloride Answers

1. Understanding the eBook Analysis Of An Unknown Chloride Answers
 - The Rise of Digital Reading Analysis Of An Unknown Chloride Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Analysis Of An Unknown Chloride Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analysis Of An Unknown Chloride Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analysis Of An Unknown Chloride Answers
 - Personalized Recommendations
 - Analysis Of An Unknown Chloride Answers User Reviews and Ratings
 - Analysis Of An Unknown Chloride Answers and Bestseller Lists

5. Accessing Analysis Of An Unknown Chloride Answers Free and Paid eBooks
 - Analysis Of An Unknown Chloride Answers Public Domain eBooks
 - Analysis Of An Unknown Chloride Answers eBook Subscription Services
 - Analysis Of An Unknown Chloride Answers Budget-Friendly Options
6. Navigating Analysis Of An Unknown Chloride Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Analysis Of An Unknown Chloride Answers Compatibility with Devices
 - Analysis Of An Unknown Chloride Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analysis Of An Unknown Chloride Answers
 - Highlighting and Note-Taking Analysis Of An Unknown Chloride Answers
 - Interactive Elements Analysis Of An Unknown Chloride Answers
8. Staying Engaged with Analysis Of An Unknown Chloride Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analysis Of An Unknown Chloride Answers
9. Balancing eBooks and Physical Books Analysis Of An Unknown Chloride Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analysis Of An Unknown Chloride Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Analysis Of An Unknown Chloride Answers
 - Setting Reading Goals Analysis Of An Unknown Chloride Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Analysis Of An Unknown Chloride Answers
 - Fact-Checking eBook Content of Analysis Of An Unknown Chloride Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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