

Chemistry Student Edition - Basic Answer Key

Chapter 11: Chemical Reactions

Chemical Equations

Questions

1. Early Chinese alchemists discovered an early form of gunpowder. What was the composition of this substance?
2. What later developments were made to the gunpowder recipe that improved its pyrotechnic properties?
3. Make an argument for why the burning of a candle is consistent with the law of conservation of matter/mass.
4. Think of an experiment that you could conduct to demonstrate that mass is conserved for a given chemical change.
5. Balance the following chemical equations:
 - a. $C + O_2 \rightarrow CO$
 - b. $CO + O_2 \rightarrow CO_2$
 - c. $H_2 + Br_2 \rightarrow HBr$
 - d. $K + H_2O \rightarrow KOH + H_2$
 - e. $O_3 \rightarrow O_2$
 - f. $N_2 + H_2 \rightarrow NH_3$
 - g. $Zn + AgCl \rightarrow ZnCl_2 + Ag$
 - h. $Cl_2 + NaI \rightarrow NaCl + I_2$
 - i. $P_4O_{10} + H_2O \rightarrow H_3PO_4$
 - j. $Be_2C + H_2O \rightarrow Be(OH)_2 + CH_4$
 - k. $S + HNO_3 \rightarrow H_2SO_4 + NO_2 + H_2O$
 - l. $NH_3 + CuO \rightarrow Cu + N_2 + H_2O$
 - m. $HCl + CaCO_3 \rightarrow CaCl_2 + H_2O + CO_2$

Answers

1. Early gunpowder was comprised of saltpeter (impure potassium nitrate), sulfur, and charcoal (carbon).
2.
 - a. Purification of potassium nitrate
 - b. Using potassium chlorate instead of potassium nitrate
 - c. Further refinement of the recipe to find the proportions of reactants that yielded the most energetic reaction
3. The burning of a candle obeys the law of conservation of mass because, even though the mass of the candle will decrease as it burns, the mass is not lost. It is transferred into carbon dioxide and water that enter the surrounding air.
4. The reaction between iron and oxygen to produce rust could be studied as an example of conservation of mass. Steel wool could be placed into a covered beaker containing water. The reaction would proceed over time and rust would begin to form. The beaker could be weighed before and after the change. The mass would be the same, even though a new

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A Loxley



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Table of Contents Chapter 11 Chemical Reactions Practice Problems Answer Keys

1. Understanding the eBook Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - The Rise of Digital Reading Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Personalized Recommendations
 - Chapter 11 Chemical Reactions Practice Problems Answer Keys User Reviews and Ratings
 - Chapter 11 Chemical Reactions Practice Problems Answer Keys and Bestseller Lists
5. Accessing Chapter 11 Chemical Reactions Practice Problems Answer Keys Free and Paid eBooks
 - Chapter 11 Chemical Reactions Practice Problems Answer Keys Public Domain eBooks
 - Chapter 11 Chemical Reactions Practice Problems Answer Keys eBook Subscription Services
 - Chapter 11 Chemical Reactions Practice Problems Answer Keys Budget-Friendly Options
6. Navigating Chapter 11 Chemical Reactions Practice Problems Answer Keys eBook Formats

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 - Chapter 11 Chemical Reactions Practice Problems Answer Keys Compatibility with Devices
 - Chapter 11 Chemical Reactions Practice Problems Answer Keys Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Highlighting and Note-Taking Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Interactive Elements Chapter 11 Chemical Reactions Practice Problems Answer Keys
 8. Staying Engaged with Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 11 Chemical Reactions Practice Problems Answer Keys
 9. Balancing eBooks and Physical Books Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 11 Chemical Reactions Practice Problems Answer Keys
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Setting Reading Goals Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Fact-Checking eBook Content of Chapter 11 Chemical Reactions Practice Problems Answer Keys
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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