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# Linear Algebra with Applications

Eighth Edition

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web jan 10 2021 equilibrium lies to the right because HCl pK<sub>a</sub> 7 is a stronger acid than NH<sub>4</sub><sup>+</sup> pK<sub>a</sub> 9.3 and NH<sub>3</sub> is a stronger base than Cl<sup>-</sup> pK<sub>b</sub> 7.7 to identify the

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web characteristics physical acids would look solid liquid or in the form of gas it would also have a sour taste bases would feel slippery and solid in nature except for ammonia

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web log H<sup>+</sup> 10<sup>-4</sup> 0.5891 x 10<sup>-5</sup> mol dm<sup>-3</sup> log pK<sub>a</sub> pH 3.75 4.05 0.30 10.0 3.0 0.50 and HCOOH 0.1 so HCOO<sup>-</sup> 0.200 mol dm<sup>-3</sup> amount vol x

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web only the better candidates could write successive ionisation equations in part c while part d about the pH of equimolar strong and weak acid solutions was poorly answered

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web a substance may be assigned to one our four conceivable categories it may be an acid or a base but in addition it may be both an acid and a base or it may be neither an acid nor

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web to detect the presence of acids and bases Michael adds a few drops of phenolphthalein indicator to two colorless solutions a and b he observes the following changes in the

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web nov 5 2023 most of the acids are solid it has a sour taste they produce hydrogen after reacting with active metals like zinc magnesium aluminium or iron acids have a low pH

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web section 2 pH of strong acids number of protons released monoprotic acid acid that releases one H<sup>+</sup> ion per molecule e.g.

hcl hydrochloric acid hno<sub>3</sub> nitric acid ch

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web jan 10 2021 s14 3 21 strong electrolytes are 100 ionized and as long as the component ions are neither weak acids nor weak bases the ionic species present result

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web problem 1 which of the following are not acids ch<sub>3</sub>cooh co<sub>2</sub> hno<sub>2</sub> hcooh cc<sub>4</sub> c h 3 c o o h c o 2 h n o 2 h c o o h c c l 4  
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web in grade 11 learners were introduced to a lot of the concepts that are expanded on in this chapter these include acid and base models and definitions conjugate acid base

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