

# REACTIVITY TRENDS IN THE PERIODIC TABLE

**Reactivity Trends in the Periodic Table**

**Reading Passage**

Read the passage carefully before answering the questions that follow.

The periodic table organizes elements based on their atomic number and properties, revealing patterns known as periodic trends. One of these trends is reactivity, which describes how likely an element is to participate in chemical reactions. Understanding these trends helps scientists predict how elements will behave in various chemical contexts.

Metals, such as sodium, calcium, and potassium, are highly reactive because they have a single valence electron that they readily lose in chemical reactions. This tendency is due to their large atomic size and low ionization energy, making it energetically favorable for them to react with other substances, especially nonmetals. Conversely, noble gases, like helium, neon, and argon, are very inert because they have full valence electron shells. This stable electronic configuration makes them unlikely to form bonds or participate in reactions under normal conditions.

Across periods in the periodic table, reactivity generally decreases as atomic number increases and valence electrons are held more tightly. In addition, atomic size influences reactivity; larger atoms tend to be more reactive if they have fewer valence electrons in their outermost shell. This electronic configuration directly impacts their ability to attract or lose electrons during chemical reactions. The periodic table also illustrates trends in ionization energy and electronegativity, which are closely related to an element's reactivity patterns. These trends, however, become more complex when considering transition metals, which may not fit as neatly into these general patterns or gain reactivity by involving their partially filled d-orbitals in chemical bonding and complex formation.

**True or False: Reactivity in the Periodic Table**

Circle whether each statement is True or False.

- Metals are highly reactive because they have single valence electrons that are difficult to lose.
- Noble gases are very reactive because they have full valence electron shells.
- Reactivity increases across a period, which generally decreases the number of valence electrons.
- Elements with fewer valence electrons tend to be less reactive than those with more valence electrons.

# Periodic Trends Reactivity Lab Answer Key

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