

Periodic Trends

ATOMIC RADIUS

1. What trend in atomic radius do you see as you go down a group/family on the periodic table?
2. What causes this trend?
3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
4. What causes this trend?
5. Circle the atom in each pair that has the largest atomic radius.
 - a) Al B
 - b) S O
 - c) Br Cl
 - d) Na Al
 - e) O F
 - f) Mg Ca
6. Put the following elements in order from smallest to largest atomic radius **and** explain why:
C, O, Sn, Sr.

ELECTRONEGATIVITY

7. Define electronegativity
8. How does the ionic radius of a nonmetal compare with its atomic radius?
9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
10. What causes this trend?
11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
12. What causes this trend?
13. Circle the atom in each pair that has the greater electronegativity.
 - a) Ca Ga
 - b) Li O
 - c) Cl S
 - d) Br As
 - e) Ba Sr
 - f) O S

GENERAL QUESTIONS

14. Which group tends to form +1 ions? _____
15. Which group tends to form +2 ions? _____
16. Which group tends to form -1 ions? _____
17. Which group tends not to form ions or react? _____

Periodic Table Chapter Test A Answer Key

Neil D. Jespersen, Pamela Kerrigan

Periodic Table Chapter Test A Answer Key:

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