

Chapter 14

- 14-1** (a) The initial pH of the NH_3 solution will be less than that for the solution containing NaOH. With the first addition of titrant, the pH of the NH_3 solution will decrease rapidly and then level off and become nearly constant throughout the middle part of the titration. In contrast, additions of standard acid to the NaOH solution will cause the pH of the NaOH solution to decrease gradually and nearly linearly until the equivalence point is approached. The equivalence point pH for the NH_3 solution will be well below 7, whereas for the NaOH solution it will be exactly 7.
- (b) Beyond the equivalence point, the pH is determined by the excess titrant. Thus, the curves become identical in this region.
- 14-2** Completeness of the reaction between the analyte and the reagent and the concentrations of the analyte and reagent.
- 14-3** The limited sensitivity of the eye to small color differences requires that there be a roughly tenfold excess of one or the other form of the indicator to be present in order for the color change to be seen. This change corresponds to a pH range of ± 1 pH unit about the pK of the indicator.
- 14-4** Temperature, ionic strength, and the presence of organic solvents and colloidal particles.
- 14-5** The standard reagents in neutralization titrations are always strong acids or strong bases because the reactions with this type of reagent are more complete than with those of their weaker counterparts. Sharper end points are the consequence of this difference.

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