

Question 35



In the early 1900s, Robert Millikan performed a series of experiments to measure the charge on tiny oil droplets suspended in an electric field. By analyzing how the droplets moved under gravity and an applied voltage, Millikan determined the fundamental electric charge (the charge on a single electron). From this charge and additional data, the mass of the electron could be calculated. In this passage, you will examine data from similar oil drop experiments to investigate the mass of the electron.



A fine mist of oil droplets is introduced into a transparent chamber situated between two horizontal metal plates. The space between the plates is observed through a microscope or viewing scope. A DC power supply applies a variable voltage across the plates, creating a uniform electric field. The droplets become charged locally through friction or ionization and can be made to rise, fall more slowly, or remain suspended depending on the balance of forces on them. The time it takes for droplets to move between reference markers and the voltage across the plates are recorded. A cross-sectional view: two parallel horizontal plates, with an oil droplet in between. A microscope is shown to the side for observing the droplet's motion. The top

In Experiment 1, why is the droplet's time to fall measured with the electric field turned off?

- A. To observe how the droplet interacts with the charged plates
- B. To determine the droplet's mass-related properties free from electrical forces
- C. To calculate the exact charge on the droplet
- D. To allow the droplet to pick up additional charge

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