

## Chapter 19 Study Guide: Acids, Bases & Titration – Answers

1. Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids **and** bases.

	ACIDS	BASES
<b>Arrhenius</b>	Form hydronium ions ( $\text{H}_3\text{O}^+$ ) in aqueous sol'n	Form hydroxide ions ( $\text{OH}^-$ ) in aqueous sol'n
<b>Brønsted-Lowry</b>	Proton ( $\text{H}^+$ ) donors	Proton ( $\text{H}^+$ ) acceptors
<b>Lewis</b>	Electron pair acceptors	Electron pair donors

2. Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:



3. Write the names or formulas for the following acids and bases:

- |  |   |
|--|---|
| a. $\text{H}_2\text{S}$ – hydrosulfuric acid       | e. Nitrous acid – $\text{HNO}_2$                |
| b. $\text{HC}_2\text{H}_3\text{O}_2$ – acetic acid | f. Calcium hydroxide – $\text{Ca}(\text{OH})_2$ |
| c. $\text{NaOH}$ – sodium hydroxide                | g. hydrochloric acid – $\text{HCl}$             |
| d. $\text{HNO}_3$ – nitric acid                    |   |

Use the following formulas for the remaining questions: (these will not be given to you on the test!)

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$M_A V_A = n_A = M_B V_B = n_B$$

4. What is the pH of  $1.34 \times 10^{-4}$  M solution of hydrochloric acid?

$$\text{pH} = -\log [1.34 \times 10^{-4} \text{M}] = 3.873$$

5. Determine the pH and pOH of a  $1.85 \times 10^{-6}$  M NaOH solution.

$$\text{pOH} = -\log [1.85 \times 10^{-6} \text{M}] = 5.733$$

$$\text{pH} = 14 - 5.73 = 8.267$$

6. What is the  $[\text{OH}^-]$  of an aqueous solution that has a pH of 5.0?

$$\text{pOH} = 9 \quad [\text{OH}^-] = 1 \times 10^{-9} \text{ M}$$

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35 \text{g NaOH}}{2.000 \text{L}} \times \frac{1 \text{mol NaOH}}{40.00 \text{g NaOH}} = 0.44 \text{ M NaOH}$$

$$\text{pOH} = -\log[0.44 \text{M}] = 0.36 \quad \text{pH} = 14 - 0.36 = 13.64 \quad \text{basic}$$

# Chapter 19 Study Guide Acids Bases Answers

**John A. Olmsted, Gregory M. Williams**

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practitioners The major textbooks are excellent references but do not allow concise reading on specific topics and are not intended to act as both text and study guide There are also several handbooks available but these are usually written for general pediatric residents and lack the advanced physiology and pathophysiology required for the higher level pediatric critical care practitioner

**Introduction to General, Organic, and Biochemistry Study Guide** William Scovell, 2006-03  
By William M Scovell This resource helps students organize their study time and guides them through the topics in a systematic way Each chapter of the text is covered by an introduction a list of review topics section by section study suggestions and questions a list of key terms and a practice exam with worked out answers

**Student Study Guide to Accompany Petrucci's General Chemistry** Robert K. Wismer, 1985 Student Study Guide to accompany Chemistry Martin Silberberg, Libby Weberg, 2005-01-06

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Solutions Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed Paul A. Bartlett, Judith G. Koch, 1992

**Student Self-study Guide** Donald W. Shive, 1975 *Human Anatomy and Physiology* John W. Hole, Jr., 1992-11

**Study Guide** George Karleskint, 1991 Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed Robert K. Wismer, 1982 *Study Guide with Answers to Selected Problems* H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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