

Chapter 18

Study Guide - Answers

1. Write the expression for the equilibrium constant for each of the following reactions.



$$K_{\text{eq}} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$



$$K_{\text{eq}} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$



$$K_{\text{eq}} = \frac{[\text{H}^+][\text{CN}^-]}{[\text{HCN}]}$$



$$K_{\text{eq}} = \frac{[\text{H}^+][\text{HSO}_4^-]}{[\text{H}_2\text{SO}_4]}$$



$$K_{\text{eq}} = [\text{Pb}^{2+}][\text{F}^-]^2$$

2. From the data provided below, calculate the value of the equilibrium constant for the reaction.



$$[\text{H}_2] = [\text{Cl}_2] = 1.0 \times 10^{-2}; [\text{HCl}] = 1.0 \times 10^{-4}$$

$$K_{\text{eq}} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$

$$K_{\text{eq}} = \frac{[1.0 \times 10^{-4}]^2}{[1.0 \times 10^{-2}][1.0 \times 10^{-2}]} = 0.0001$$



$$[\text{N}_2] = 4.4 \times 10^{-2}; [\text{H}_2] = 1.2 \times 10^{-3}; [\text{NH}_3] = 3.4 \times 10^{-3}$$

$$K_{\text{eq}} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$

$$K_{\text{eq}} = \frac{[3.4 \times 10^{-3}]^2}{[4.4 \times 10^{-2}][1.2 \times 10^{-3}]^3} = 0.152$$

3. For the reaction $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$, $K_{\text{eq}} = 0.2$. At a particular time, the following concentrations are measured: $[\text{N}_2\text{O}_4] = 2.0 \text{ M}$, $[\text{NO}_2] = 0.2 \text{ M}$. Is this reaction at equilibrium? If not, in which direction will the reaction proceed?



$$[\text{N}_2\text{O}_4] = 2.0 \text{ M}$$

$$[\text{NO}_2] = 0.2 \text{ M}$$

$$Q = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$$

$$Q = \frac{[0.2]^2}{[2.0]} = 0.02$$

$$K_{\text{eq}} = 0.2$$

$Q < K_{\text{eq}}$, so the reaction will proceed to the right (forward).

Chapter 18 Study For Content Mastery Chemical Equilibrium

Albert A Gayle



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Principles of Chemical Equilibrium Kelso B. Morris,1965 [Chemical Equilibrium In a Nutshell](#) Brian Coleman,2018-06-25 Understanding the math and minutiae of chemical equilibrium can be a tall task for anyone so why not enlist the help of a scientific squirrel to guide you on your journey Join Dr Wash as we dabble in equilibrium constants and other tools needed to predict chemical processes This book focuses on introductory concepts at the high school and early university level focusing on identifying equilibrium calculating K and Q discussing Le Chatelier s principle and tying equilibrium with the field of Thermodynamics Full of step by step instructions and practice questions this book aims to simplify one of the more complex topics found within the field of chemistry [Chemical Equilibrium and Solutions](#) James E. Banks,1967 [Improving the Understanding of Chemical Equilibrium in Second Semester General Chemistry Students](#) David C. Swanson,California State University, Sacramento,2011 The concept of chemical equilibrium is central to the understanding of chemistry and is widely viewed as the most difficult topic to master in the General Chemistry curriculum The study of chemical equilibrium is at the core of Chem 1B the second semester general chemistry class at California State University Sacramento This study attempts to locate the primary reasons for students difficulties with chemical equilibrium and to suggest help for their problems particularly with regard to the conceptual understanding of this topic The action research technique has been employed to characterize this problem develop and introduce interventions to the student environment meant to improve the problem and collate and summarize the results of this intervention It is found that there is a lack of conceptual understanding of the principles underlying chemical equilibrium on the part of general chemistry students and it is recommended that additional conceptual teaching be added to the general chemistry curriculum at CSUS

Chemistry ,2001-04-01 *Chemical Equilibrium* P. H. Carnell,R. N. Reusch,1962 [Workbook 15](#) Thobeka Pete,Turret Correspondence College (Johannesburg),1988 **Freshman Chemistry Problems and how to Solve Them** Michell J. Sienko,1964 **Fundamental Chemical Equilibria** Kelso B. Morris,1971 *Chemical Equilibria in Solution* M. J. Blandamer,1992

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