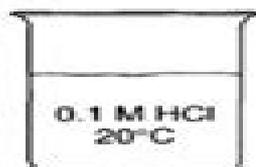
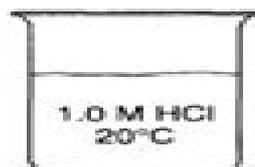


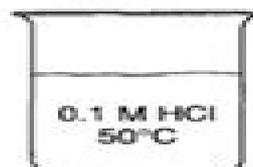
- A reaction is most likely to occur when reactant particles collide with
 - proper energy, only
 - proper orientation, only
 - both proper energy and proper orientation**
 - neither proper energy nor proper orientation
- The energy needed to start a chemical reaction is called
 - potential energy
 - kinetic energy
 - activation energy**
 - ionization energy
- After being ignited in a Bunsen burner flame, a piece of magnesium ribbon burns brightly, giving off heat and light. In this situation, the Bunsen burner flame provides
 - ionization energy
 - activation energy**
 - heat of reaction
 - heat of vaporization
- As the temperature increases, the rate of an exothermic reaction
 - decreases
 - increases**
 - remains the same
- Which conditions will increase the rate of a chemical reaction?
 - decreased temperature and decreased concentration of reactants
 - decreased temperature and increased concentration of reactants
 - increased temperature and decreased concentration of reactants
 - increased temperature and increased concentration of reactants**
- Base your answer to the following question on In each of the four beakers shown below, a 2.0-centimeter strip of magnesium ribbon reacts with 100 milliliters of $\text{HCl}(\text{aq})$ under the conditions shown.



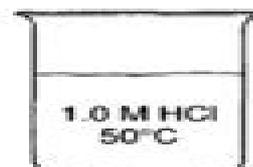
Beaker A



Beaker B



Beaker C



Beaker D

In which beaker will the reaction occur at the fastest rate?

- A
- B
- C
- D**

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